

Doped Organic Transistors

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Abstract

Organic field-effect transistors hold the promise of enabling low-cost and flexible electronics. Following its success in organic optoelectronics, the organic doping technology is also used increasingly in organic field-effect transistors. Doping not only increases device performance, but it also provides a way to fine-control the transistor behavior, to develop new transistor concepts, and even improve the stability of organic transistors.

This review summarizes the latest progress made in the understanding of the doping technology and its application to organic transistors. It presents the most successful doping models and an overview of the wide variety of materials used as dopants. Further, the influence of doping on charge transport in the most relevant polycrystalline organic semiconductors is reviewed and a concise overview on the

influence of doping on transistor behavior and performance is given. In particular, recent progress in the understanding of contact doping and channel doping is summarized.

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1 Introduction

The advent of microelectronics has changed our everyday life. The key element is the field effect transistor, which has been introduced in silicon in the early 1960s. In the city of Dresden, the largest microelectronics center in Europe, more than a million transistors are produced everyday - per human being on earth! Current microelectronics technology is almost exclusively based on the technology of single crystalline silicon. While the technology has been scaled over many orders of magnitude to reach highest performance in speed, cost, and power reduction, there are many novel applications where this technology is less suited since properties like large-area, flexibility and easy processing e.g. by printing would be required.

Here, organic electronics might come into play as a new paradigm: Organic electronics is based almost entirely on carbon, can be deposited by simple fabrication at low temperatures on almost any substrate, including flexible and transparent materials, and offers a wide variety of functions, from switching over memory, light emission and absorption, to sensorics. The main drawback is that organic semiconductors are characterized by much lower mobilities than silicon and other crystalline inorganic semiconductors, thus limiting speed and current driving possibilities. Due to these properties, organic electronics will not compete with, but complement silicon-based electronics in the future.

The key devices like organic field-effect transistor (OFET), organic light-emitting diode (OLED) and organic solar cell (OSC) were all first presented in the 1990s. At the time when this article is written, OLED have already reached a multi-billion US\$ market as

small displays, mainly used in smart phone displays and are rapidly entering the TV market. Lighting based on white OLED and OSC are making the first steps into markets, while organic transistors have so far not reached the performance required for broad application in products, although they might have the highest market potential of all organic electronics devices.

When comparing the development of organic electronics with classical silicon electronics, one surprising observation is that *doping*, i.e. the addition of impurities to control the Fermi level and raise the conductivity, was initially almost completely disregarded in organics. Typically, organic materials have rather large impurity concentrations which, in association of usually larger gaps on the order of 2 eV, make it highly unlikely that they are intrinsic and that the Fermi level is positioned in the center of the gap. Nevertheless, there are only few publications which consider the doping of organics from the early days of organic electronics. For instance, Yamamoto et al. used oxidizing gases such as iodine or bromine vapor to dope phthalocyanines.¹ Although this leads to a doping effect, the films such prepared are not stable.

The same material was also doped with organic acceptor molecules such as ortho-chloranil,² tetracyano-quinodimethane (TCNQ) or dicyano-dichloro-quinone (DDQ),^{3,4} which produces more stable doping. Covalently bound stacked phthalocyanines⁵ and oligothiophenes⁶ were doped by DDQ, which is a rather effective dopant but has low thermal stability.

The importance of doping became obvious when highly efficient and stable doping was reported,⁷ combined with the application in OLEDs,⁸ where the p-doped hole transport layers led to significantly reduced operating voltages. Since then, the technique has been widely extended and refined and is today in broad commercial application as almost all commercially available OLED displays contain transport layers which employ controlled doping. Similarly, dopants are applied in highly efficient white OLED for lighting applications⁹ and OSCs, where the most efficient devices presented so far use

doping techniques.^{10,11}

The main effects which have made doping successfully used in organic devices are twofold:

- First, doped layers have much higher conductivity, which reduces bulk ohmic losses and allows thicker transport layers important for optical optimization and shortcut avoidance
- Second, doped layers create thin space charge layers at contacts, allowing efficient contacts even in case of electronic barriers due to work function mismatch¹²

Naturally, one might ask whether the positive effects of doping also apply to the organic device class, which showed the least impact so far, organic transistors. Indeed, it has been shown that doping is improving performance of organic transistors (see e.g.^{13–20}), can be used to fine-tune transistor behavior,^{21–24} and increases the stability of OFETs.^{25,26}

In this review, we discuss the recent progress of the application of doping in organic transistors in detail. After a short summary of different doping models, we discuss the chemistry and synthesis of dopants. Furthermore, we address the doping of polycrystalline materials with higher mobility and finally the device application in doped organic transistors.

2 Physics of Doping

Organic doping resembles doping in inorganic semiconductors in many aspects. An impurity is added to a host material, which either donates an electron to the host (n-doping) or accepts an electron from the host, leaving a hole behind (p-doping).

Doped organic films can either be processed by co-evaporation of host and dopant in vacuum,⁷ by adding a defined impurity into a solution of the host semiconductor and

process the film wet-chemically, e.g. by spin-coating,²⁷ or by exposing the semiconducting film or crystal to a gas of the dopant.¹

Despite the ease of processing, the details of the organic doping process are surprisingly complex and still not understood in its entirety. Only recently, significant progress has been made that allowed to describe doping consistently over a wide range of doping concentrations. In the following, this progress is summarized, which will provide a starting point for the discussion of doping in organic transistors.

2.1 Thermal Activation of Organic Doping

In the most intuitive understanding of doping, an integer number of electrons is transferred from the host to the dopant (p-doping) or from the dopant to the host (n-doping).^{7,28-33} In order to make this transfer efficient, the electron affinity (EA) of the dopant has to exceed the ionization potential (IP) of the host (p-doping) or the IP of the dopant has to be smaller than the EA of the host (n-doping) (see Figure 1). It has to be noted that the IE and EA have to be measured in the solid state, which are usually smaller than the energy levels in the gas phase.

However, this simplified understanding of doping does not fully capture the doping process. For example, although the EA of the popular p-dopant 2,2'-(perfluoronaphthalene-2,6-diylidene)dimalononitrile (F₆-TCNNQ) (EA \approx 5.07 eV) and the IP of the hole transport material N,N,N',N'-tetrakis(4-methoxyphenyl)-benzidine (MeO-TPD) (IP \approx 5 eV) almost match, electron transfer between host and dopant is not spontaneous, but shows a significant activation energy of $E_A \approx$ 0.46 eV,³⁵ i.e. additional energy has to be supplied to generate a free hole in the pentacene matrix.

Several explanations have been put forward to explain the origin of the activation energy. Often, it is argued that the activation energy reflects the low permittivity of the organic semiconductors, which leads to large Coulomb interactions between charge carriers and ionized dopant molecules.^{36,37} Even if electron transfer between host and

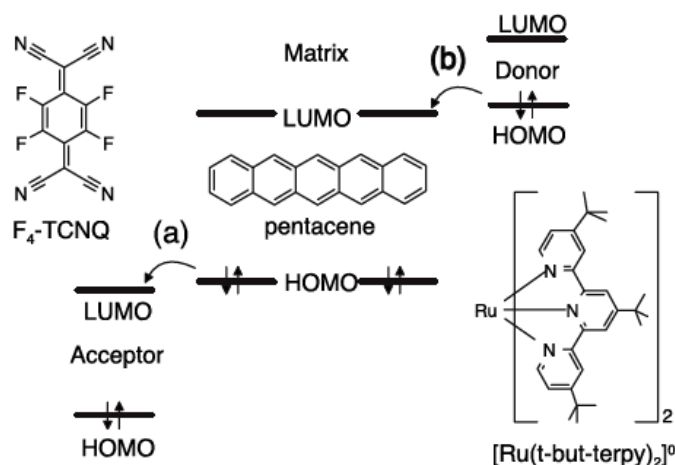


Figure 1: Simplified model of the organic doping process exemplified by the organic semiconductor pentacene doped with F₄-TCNQ (p-doping) or [Ru(t-but-terpy)₂]⁰ (n-doping). Free charge carriers are generated by an electron transfer between dopant and matrix molecule. To facilitate electron transfer, the LUMO of the p-dopant has to be lower than the HOMO of the matrix molecule (for p-doping, left) or the HOMO of the n-dopant has to exceed the LUMO of the matrix molecule (for n-doping, right). Reprinted from³⁴. Copyright 2008 American Physical Society.

dopant is efficient, the transferred charge carrier still experiences a strong electrostatic attraction to the ionized dopant molecule. A charge transfer complex is formed between the host and the dopant, which has to be dissociated before a *free* electron or a *free* hole is generated. In fact, this understanding of the doping process is corroborated by scanning tunneling microscopy studies of Ha et al.,³⁸ who studied doping of pentacene by the p-dopant tetrafluoro-tetracyano-quinodimethane (F₄-TCNQ). They observe a cloud of positive charge around ionized dopant molecules, which they interpret in terms of a localization of the donated hole by the Coulomb potential of the ionized dopant.

The electrostatic attraction between the ionized dopant molecule and the hole or electron transferred to a neighboring molecule can be significant. Mityashin et al. even argued that in perfect organic crystals, the binding energy of this charge pair is too large to be overcome by thermal excitation, which means that it is highly improbable that free charge carriers are generated at all.³⁶ To explain doping observed at room

temperature, they showed that the dissociation probability of the charge pair is increased due to increased disorder in the energy landscape at higher doping concentrations. Due to increasing disorder, more and more percolation paths are opened along which the hole on the matrix can escape the Coulomb attraction of the dopant.

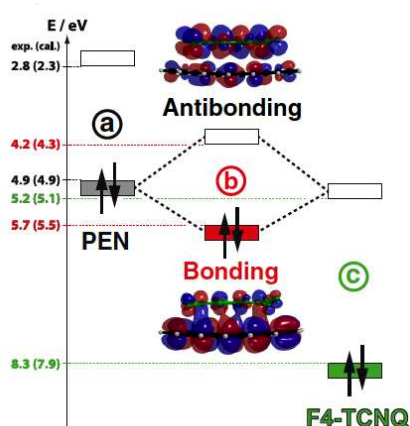


Figure 2: Doping via hybrid states. Instead of assuming an integer charge transfer between dopant and matrix, some authors argue that a hybrid state is formed by coupling the HOMO of the matrix and the LUMO of the dopant. The activation energy observed in experiments is reflected by the difference between the antibonding orbital of the hybrid state and the HOMO state of the host. In the figure, the model is sketched for the example of doping Pentacene by F₄-TCNQ. Reprinted with permission from³⁹. Copyright 2012 American Physical Society.

A slightly different approach to explain doping was proposed recently by the Koch group.³⁹⁻⁴¹ It is argued that the simplified model of doping shown in Figure 1 implies that an integer number of charges are transferred between dopant and matrix, which might not be the case for all host and dopant combinations. In fact, charge transfer complexes formed during doping can be held together by both ionic and covalent contributions.⁴² Integer charge transfer is assumed to take place if charge transfer complexes have mostly ionic character, which is often observed if the donor highest occupied molecular orbital (HOMO) is substantially higher in energy than the acceptor lowest unoccupied molecular orbital (LUMO).^{43,44}

In contrast, charge transfer states arising from interactions where the donor HOMO

is near or lower in energy than the acceptor LUMO will have a more covalent character and will involve significant hybridization of the donor and acceptor orbitals.⁴¹ Salzmann et al.³⁹ proposed that these hybrid states formed by coupling the HOMO of the matrix and the LUMO of the dopant (for p-doping) are dominating doping. A bonding and an antibonding state are formed (see Figure 2). While the bonding state of the hybrid is fully occupied by electrons, the antibonding state is located in the energy gap of the matrix and almost empty. However, due to thermal excitation some electrons can be excited from the HOMO of the matrix to the antibonding state of the hybrid leaving behind a hole on the matrix.⁴⁵

Assuming partial charge transfer and the formation of a hybrid dopant/matrix complex, the activation energy for the doping process is given by the difference between the antibonding state of the hybrid and the HOMO of the matrix (for p-doping). This activation energy E_A can be approximated as^{40,41}

$$E_A = \frac{H_M - L_D}{2} \left[1 + \sqrt{1 + \frac{4\beta^2}{(H_M - L_D)^2}} \right] \quad (1)$$

with H_M : HOMO energy of the matrix, L_D : LUMO energy of the dopant, and β : resonance integral between the HOMO of the matrix and the LUMO of the dopant.

2.2 Charge Carrier Statistics in Doped Organic Layers

Regardless of the microscopic details of the doping process and the origin of its activation energy, a straightforward statistical model capable of describing the doping process over a wide range of doping concentrations and for different host/matrix was proposed recently.^{35,46,47}

In the model, the charge neutrality equation for a p-doped semiconductor is solved self-consistently. To obtain a good agreement to experiments, trap states had to be included. The charge neutrality equation reads:

$$p + N_T^+ = n + N_A^- \approx N_A^- \quad (2)$$

with p : density of free holes, N_T^+ : density of positively charged traps, n : density of electrons, N_A^- : density of ionized dopant molecules.

The density of free holes, p , and the density of charged trap states, N_T^+ obey the Fermi-Dirac integrals

$$p = \int_{-\infty}^{\infty} dE g_{DOS}(E) [1 - f(E)] \quad (3)$$

$$N_T^+ = \int_{-\infty}^{\infty} dE g_T(E) [1 - f(E)] \quad (4)$$

Here, $f(E)$ is the Fermi-Dirac distribution and $g_{DOS}(E)$ the density of mobile states in the HOMO of the matrix and $g_T(E)$ the density of trap states. A gaussian distribution g_{DOS} is assumed, sometimes modified by an exponential tail.^{35,46}

Finally, the doping process itself is described by thermal activation of the doped charge carrier from a discrete doping level at E_A toward an effective transport level in the HOMO density of states (DOS). The density of ionized dopants N_A^- thus reads

$$N_A^- = \frac{N_A}{1 + \exp[(E_A - E_F)/k_B T]} \quad (5)$$

with N_A density of dopant molecules in the film, E_F Fermi level, k_B Boltzmann constant, and T temperature.

Representative results of this model are shown as solid lines in Fig. 3. The model is used to determine the Fermi level position in the amorphous organic semiconductor MeO-TPD, which is a standard hole transport material,⁴⁸ doped by the p-dopant fluorinated fullerene $C_{60}F_{36}$ ^{49,50} at various doping concentrations.

In Fig. 3 the calculated Fermi level position is compared to the experimental Fermi level position obtained by photoelectron spectroscopy^{35,46} (symbols). The Fermi level is

plotted with respect to the HOMO position of MeO-TPD, i.e. the HOMO position is set to 0 eV. Different purities of MeO-TPD are used: as received (red), purified twice (black) or four times (blue) by sublimation. Overall, an excellent agreement between the model and experimental data is found. Similar calculations were as well reported by Salzman et al. in 2015 for quaterthiophene doped by F₄-TCNQ.⁴⁵

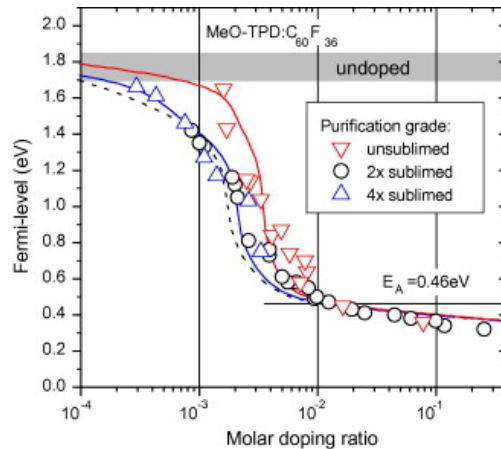


Figure 3: Shift of the Fermi level in the hole transport material MeO-TPD upon doping by C₆₀F₃₆. Different grades of MeO-TPD are compared (used as received, purified by two sublimation cycles, purified by 4 sublimation cycles). The model calculations (lines) fit the measurements (open symbols) well. Reprinted with permission from⁴⁶. Copyright 2013 Elsevier.

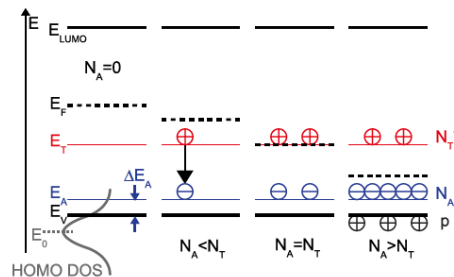


Figure 4: Simplified model to explain the complex shift of Fermi level position in MeO-TPD observed in Figure 3. At low concentrations, all donated holes are trapped by charge traps in MeO-TPD and no free charge carriers are generated. Once all traps are filled, the Fermi level shifts rapidly toward the HOMO. Reprinted from³⁵. Copyright 2012 American Physical Society.

The numerical model can be used to study the details of the doping process. The result of these studies are sketched in Figure 4. At low doping concentrations, doping only fills

trap states in the organic semiconductor and no free charge carriers are generated^{35,51,52} (see Figure 4 for $N_A < N_T$). Consequently, the Fermi Level is pinned at the trap level. Once the doping concentration exceeds the trap density (see Figure 4 for $N_A > N_T$), free holes are generated in the HOMO of MeO-TPD and the Fermi level shifts rapidly toward the HOMO at 0 eV. As expected, complete filling of all traps and the rapid shift of the Fermi level toward the HOMO occurs earlier for highly purified materials (blue and black line in Fig. 3) as in materials with a lower purity (red line), indicating the lower trap density in purified MeO-TPD.

Finally, the shift of the Fermi level saturates at high doping concentrations. Again, this behavior can be understood by the detailed charge carrier statistics as described by the above equations. The HOMO of the organic matrix material is approximated by a gaussian DOS. Therefore, there is a significant DOS tailing into the gap of the matrix material, which pins the Fermi level close to the transport states in the HOMO.

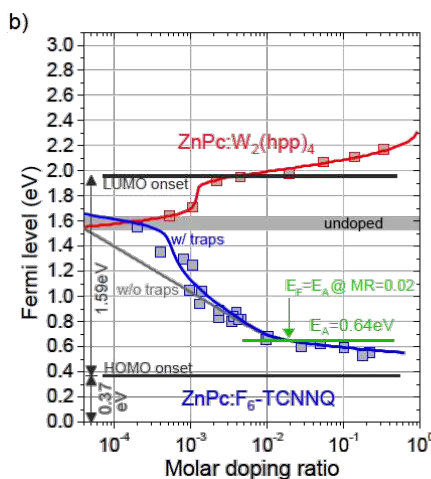


Figure 5: P- and n-type doping of the polycrystalline material ZnPc. The model discussed in Section 2.2 (solid lines) reproduces the experimental data (squares) well. Reprinted from⁴⁷. Copyright 2015 John Wiley and Sons.

The same model can be used to study doping in polycrystalline materials as well. For example, in Figure 5 the Fermi level shift in the polycrystalline material ZnPc doped by F₆-TCNNQ (p-doping) or tetrakis(1,3,4,6,7,8-hexahydro-2H-pyrimido[1,2-

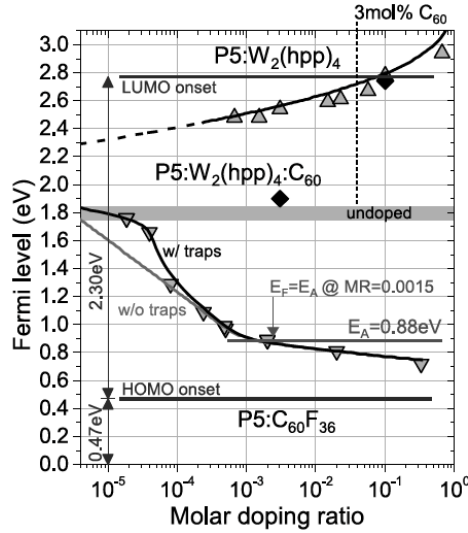


Figure 6: Shift of the Fermi level in pentacene upon p- and n-type doping. Symbols denote UPS measurements, whereas the lines are obtained by calculations. Different doping regimes are observed for p-doping: trap-limited regime at low concentrations, saturation regime at intermediate concentrations, and impurity reserve regime at large doping concentrations. The shift of the Fermi level upon n-doping is stronger and does not saturate close to the LUMO energy. Reprinted from⁴⁷. Copyright 2015 John Wiley and Sons.

a]pyrimidinato)ditungsten (II) ($W_2(hpp)_4$) (n-doping) is plotted versus the doping concentration. Furthermore, the shift of the Fermi level in pentacene upon p-doping by $C_{60}F_{36}$ and n-type doping by $W_2(hpp)_4$ is shown in Figure 6. In both cases, the experimental data obtained by UPS (symbols) agrees well with the model calculation (solid lines).

Several different doping regimes can be identified for p-doping ZnPc and pentacene. Similar to the results of the amorphous material MeO-TPD discussed above (Figures 3 and 4), doping fills almost exclusively traps at low concentrations (below 4×10^{-5} for Pentacene) and the Fermi level is pinned close to the mid-bandgap position (trap-limited regime). At medium concentrations ($8 \times 10^{-5} < c < 10^{-3}$ for pentacene) most traps are filled and the Fermi level shifts rapidly toward the HOMO. Still, the Fermi level is located above the activation energy of the dopant and, as will be discussed in the next section, most dopants are activated. This regime therefore resembles the saturation regime observed in

inorganic semiconductors.⁴⁷

At a molar doping ratio of 1.5×10^{-3} the Fermi level finally crosses the activation level of the dopant before its position finally saturates close to the HOMO. Thus, at these high doping concentrations the Fermi level is situated between the dopant level and the HOMO. Statistically, not all dopants are activated, which is known as impurity reserve regime in inorganic semiconductors.

2.3 Doping Efficiency

In recent publications, the doping efficiency (i.e. the ratio of the number of free charge carriers to the number of dopant molecules) was discussed controversially. Although early observations showed a nearly full transfer of charges from the dopant to the matrix molecule,³² this does not necessarily mean that these charge carriers are free and contribute to the current flow. In later experiments the number of free charge carriers and the actual doping efficiency was determined.^{27,53–60} Surprisingly, in most reports the doping efficiency is only on the order of 10% (16% for an n-doped system,⁵³ 4% for p-type doping using F₄-TCNQ,⁵⁴ and less than 2% for oxide based dopants (molybdenum trioxide (MoO₃))⁵⁸). However, some reports show a higher doping efficiency,⁵⁶ in particular at very low doping concentrations.⁶¹

Several explanations for the relatively low doping efficiency have been put forward. Clustering of dopant molecules was observed for transition metal oxides,^{30,55,58} and it was argued that the doping efficiency is given by a product of the dispersion efficiency and the efficiency of forming the dopant/matrix pair to form a charge transfer complex.^{62,63} Furthermore, the doping efficiency can be limited by intrinsic charge carrier traps as discussed above,^{46,52} or, as discussed above as well, the large dopant activation energy⁶³ caused by either strong Coulomb interaction^{35,36} or a strong hybridization of dopant and matrix molecule.⁴⁰

Another explanation adding to the limited doping efficiency was proposed recently.⁴⁷

It was pointed out that doped organic semiconductors usually show a significant activation energy and furthermore relatively high doping concentrations (several wt.%) are used in most organic devices. Hence, as discussed above, organic doping is used in the "freeze-out" or impurity reserve regime, where only a fraction of available dopants are thermally activated.⁴⁷ If the doping concentration is reduced, i.e. if the semiconductor is studied in the saturation regime, higher doping efficiencies can be obtained. For example, Pahner et al. were able to show that the p-doping efficiency in pentacene is very high and reaches almost unity (Figure 7).⁶¹ However, the doping efficiency drops rapidly for increasing doping concentrations, i.e. if the semiconductor is operated in the impurity reserve regime.

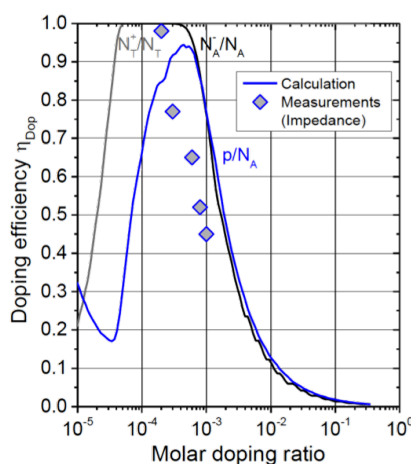


Figure 7: Efficiency of p-doping in pentacene. At low concentrations the doping efficiency increases strongly to almost 100%. Reprinted from⁴⁷. Copyright 2015 John Wiley and Sons.

3 Chemistry of Dopants

This section will briefly survey and compare the major classes of p- and n-dopants commonly used in organic semiconductors. For both the p- and n-doping sections the classification of dopants was driven in part by mechanistic or structural similarities, but

it should also be recognized that this classification was predominantly used to simplify the discussion of an extremely diverse group of chemicals. There are many distinctions in the reaction pathways possible for each of the dopants within a class and even for the same dopant with different hosts.

3.1 P-Dopants

P-dopants can be classified into several groups including: small molecules, polyelectrolytes, covalent solids, Brønsted and Lewis acids, and elemental species. Undoubtedly, some dopants could be categorized into more than one of these groups, but for the purpose of this section these groupings will be employed.

3.1.1 Elemental Species

The following diatomics have all been recognized as p-dopants for organic semiconductors: I_2 , Br_2 , Cl_2 , and O_2 .⁶⁴ The doping of polyacetylene by chlorine, bromine, and iodine is an early example of p-doping a conjugated polymer.⁶⁵ In addition to polyacetylene doping, iodine has been used to dope pentacene, phthalocyanines, and polythiophenes, among others.^{66,67} The halogens, especially iodine, can form polyanions which further complicates their interaction with hosts.⁶⁶

Oxygen is predominantly introduced as a contaminant upon incidental exposure to air. Exposure to oxygen is known to p-dope some organic semiconductors.⁶⁸ At least in the case of polythiophenes it is plausible that oxygen doping occurs by a dative bond between sulfur lone pairs and the dioxygen singly occupied molecular orbital (SOMO), or by photoinduced doping from singlet oxygen.^{69,70}

3.1.2 Covalent Solids

Examples of metal oxide p-dopants include MoO_3 , V_2O_5 , WO_3 , ReO_3 , and Fe_3O_4 .^{59,71-77}

The electron affinity of metal oxide dopants depends on the processing conditions employed, and it is well-known that their electron affinity can decrease by as much as 1 eV upon air-exposure.⁷⁸ The most commonly used metal oxide dopant is MoO₃ because it is evaporated at relatively low temperatures and can be deposited from a crucible ($\approx 400^\circ\text{C}$). The evaporation temperatures of V₂O₅ and WO₃ are much higher which leads to oxygen release and substantial heating of the substrate to be coated. Sol-gel processing of metal oxides can be used to circumvent the high deposition temperatures. Depending on the processing method used, metal oxide clusters have highly heterogeneous structures and sizes which can influence the charge generation efficiency.⁵⁵ Metal oxide dopants and their uses and processing methods in organic electronics have recently been reviewed.⁷⁸

3.1.3 Brønsted and Lewis Acids

Practically all of the common inorganic Brønsted acids have been used as dopants for poly(anilines) (PANI) and poly(pyrroles). The doping of PANI-related polymers has been reviewed previously.⁷⁹ Recently, Brønsted acid doping has been invoked to explain the surface doping of pentacene by acidic monolayers.⁸⁰ It was proposed that pentacene near the substrate-pentacene interface is protonated to form holes resulting in a threshold voltage shift (see Figure 8). Proton transfer doping has also been used to explain the doping of polythiophenes by tridecafluoro-1,1,2,2-tetrahydrooctyl trichlorosilane (FTS).⁸¹ Additionally, depending on the pH, proton transfer doping impacts the synthesis of poly(3,4-ethylenedioxythiophene):poly (4-styrenesulfonate) (PEDOT:PSS) PEDOT:PSS and other related conducting polymers.

The most common inorganic Lewis acid dopants are FeCl₃ and SbCl₅ which have both been used to dope carbon nanotubes (CNTs) and N,N'-diphenyl-N,N'-bis(3-methylphenyl)-1,1'-biphenyl-4,4'-diamine (TPD) among other materials.^{82,83} Lithium bis(trifluoromethylsulfonyl)imide (TFSI) has found widespread use as a p-dopant for 2,2',7,7'-tetrakis(N,N-di-*p*-methoxyphenyl-amine)-9,9'-spiro-bifluorene (spiro-MeO-TAD);

though, this likely generates the actual dopant spiro-MeO-TAD-TFSI from molecular oxygen.^{84,85} Bismuth carboxylates have also been used as p-dopants, and generally show a stronger doping effect with increasing fluorination of the ligand.⁸⁶ Similarly, Cu, Rh, Mo, and Cr carboxylates have been used as p-dopants with trifluoroacetate or fluorinated benzoate ligands.^{86,87}

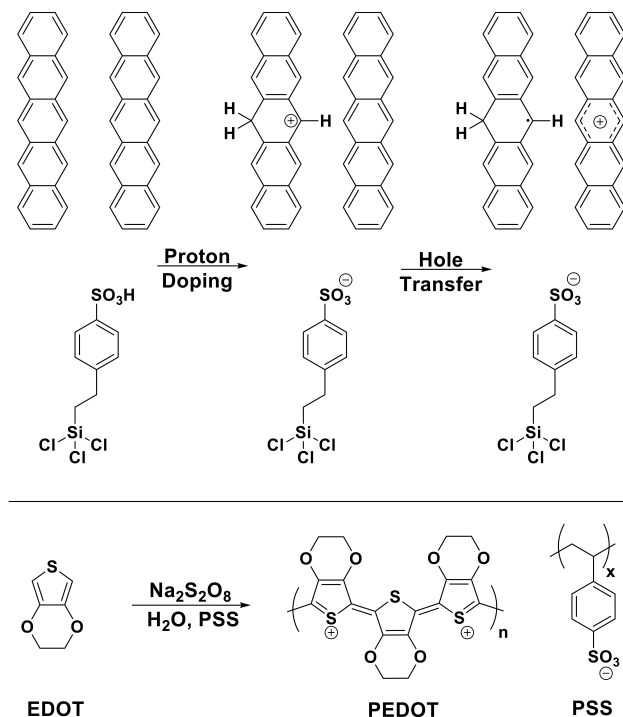


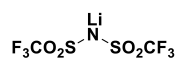
Figure 8: (Top) Brønsted acid doping of acenes by acidic monolayers.⁸⁰ (Bottom) Synthesis and structure of PEDOT:PSS.⁸⁸

3.1.4 Small Molecules

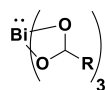
The most common molecular p-dopant is $\text{F}_4\text{-TCNQ}$.⁸⁹ This material is a strong electron acceptor with a solid state electron affinity of 5.2 eV, and the planarity of $\text{F}_4\text{-TCNQ}$ enables efficient blending with many hosts.^{90,91}

Some related TCNQ compounds form charge transfer complexes of remarkable structural order with donor molecules such as tetrathiafulvalene (TTF).⁹² Derivatives

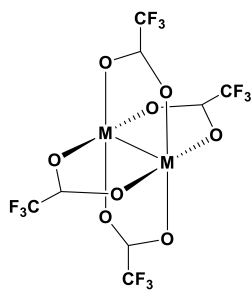
Lewis Acidic p-Dopants:



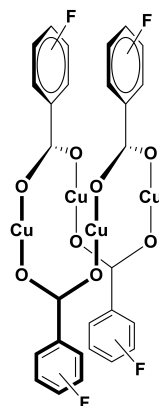
LiTFSI



Bismuth Carboxylates
R = Me, CF₃, and various
bis(trifluoromethyl)benzoates,
and 4-perfluorotoluates



M₂TFA₄ Complexes
M = Ru, Rh, Cu, Mo, Cr



Fluorinated Copper
Benzoates

Figure 9: Lewis acidic p-dopants. ^{84,86,87}

of F₄-TCNQ include 2-(3-(adamantan-1-yl)propyl)-3,5,6-trifluoro-7,7,8,8-tetracyanoquinodimethane (F₃-TCNQ-Ad1), 3,6-difluoro-2,5,7,7,8,8-hexacyanoquinodimethane (F₂-HCNQ), and F₆-TCNNQ. The electron affinity of F₄-TCNQ decreases slightly upon replacement of each fluorine with a hydrogen atom from 5.2 eV for F₄-TCNQ to 4.2 eV for TCNQ.^{35,41,93,94} F₃-TCNQ-Ad1 was developed to enable solution processing and to limit the diffusivity of this molecule in organic layers.⁹³

Much larger fluorinated organic molecules based of fullerenes have also been used as organic dopants, most notably C₆₀F₃₆⁵⁰ and C₆₀F₄₈.⁹⁵ The high molecular weight of these dopants enables extremely low doping concentrations to be achieved by standard vacuum processing.^{47,61} The Marder and Kahn groups have recently introduced molybdenum tris[1,2-bis(trifluoromethyl)ethane-1,2-dithiolene] (Mo(tfd)₃) (EA = 5.6 eV) and a more soluble analog as very potent p-dopants.^{74,96,97}

3.2 N-Dopants

The low ionization energies of highly reducing n-dopants and host radical anions make air-stable n-dopants and n-doped materials the exception to the rule. Accordingly, much of the work to develop new n-dopants has focused on enhancing the air-stability of either the precursor dopant or the doped charge transfer state. From this research several types of n-dopants have emerged including: elemental dopants, inorganic solids, low ionization potential donor molecules, organic salts, anionic dopants, hydride donor molecules, and dimers.

3.2.1 Elemental Dopants

The first and most fundamental class of n-dopants are the alkali and alkaline earth metals. Li, Na, K, and Cs are all known to reduce a wide range organic semiconductors including materials with LUMOs well above -3 eV.⁹⁸⁻¹⁰¹ As with p-type elemental dopants, alkali cations diffuse readily through organic layers leading to short operational lifetimes.⁹⁸

Small Molecule p-Dopants

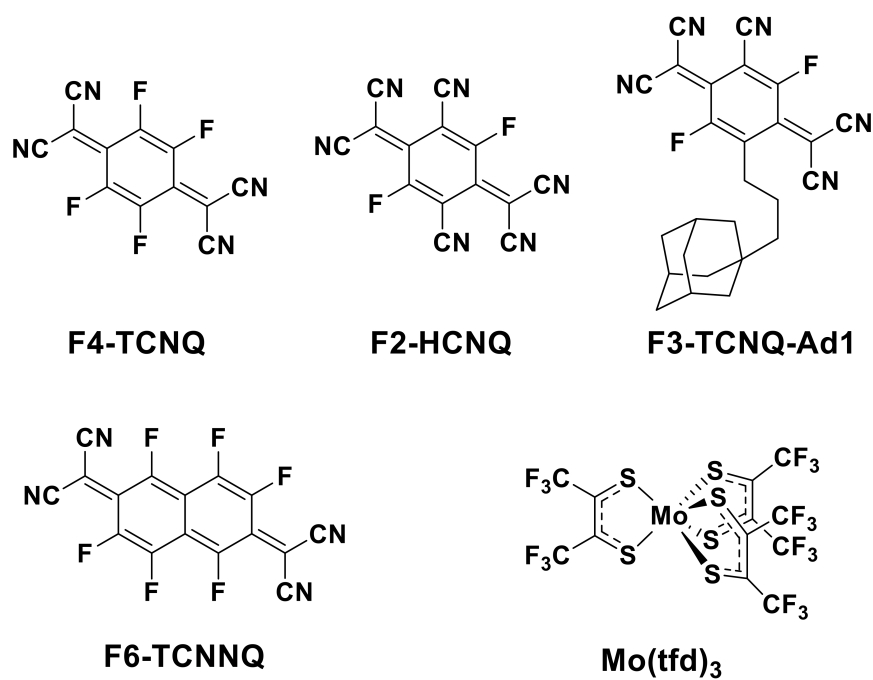


Figure 10: Examples of high electron affinity molecular p-dopants. [35](#), [41](#), [89](#), [93](#), [94](#), [96](#)

Moreover, the alkali metals are themselves difficult and often dangerous to handle.

3.2.2 Inorganic Salts

A much more easily handled class of dopants are the oxy metal salts of the general form X_2CO_3 where $X = Li, Na, K, Rb, Cs$.^{102–106} Cs_3PO_4 , Cs_3VO_4 , CsN_3 , and Li_3N have also been used as n-dopants.^{86,106–108}

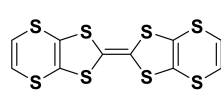
In all cases, these salts are processed via high temperature vacuum deposition, and at present it is not clear what products are responsible for their doping effect due to the formation of various oxides and suboxides.^{109,110}

3.2.3 Low Ionization Potential Donors

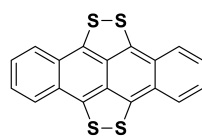
N-dopants with low ionization energies are intrinsically sensitive to oxidation, and thus care must be taken in handling these materials. Several n-dopants reported in the literature are summarized in Figure 11. TTF and the derivative bis(ethylenedithio) (BEDT)-TTF have been used as n-dopants for TCNQ, phthalocyanines, and naphthalenetetracarboxylic dianhydride (NTCDA).^{92,111} Previous reviews have enumerated the various TTF derivatives and their corresponding electrochemistry.¹¹² The neutral tris[4-(dimethylamino)phenyl]methyl radical ($E^{0/-} = -1.16$ V vs Fc/Fc^+) has been used as a solution processed dopant for the tris(4-nitrophenyl)methyl radical.¹¹³ The donor TTN has been used as an n-dopant for hexadecafluorophthalocyaninatozinc (F_{16} -ZnPc).¹¹⁴

Bloom and Leo have also used bis-(terpyridine)ruthenium(II) $[Ru(terpy)_2]^0$ as an n-dopant for ZnPc which has a substantially higher LUMO than F_{16} -ZnPc making it even more difficult to dope (dopant = -1.7 V vs Ag/Ag^+).¹²³ Marder and Kahn have demonstrated potent n-doping by cobaltocene and decamethylcobaltocene which have solid state ionization energies of 4.07 eV and 3.30 eV, respectively.^{71,118,124,125} The donor with perhaps the lowest ionization energy for a molecular dopant is $W_2(hpp)_4$ which has a solid state ionization energy of only 2.40 eV.^{121,122,126}

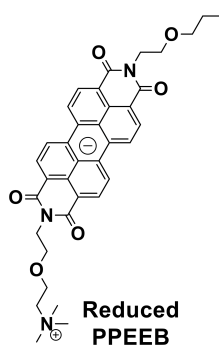
Low Ionization Potential n-Dopants



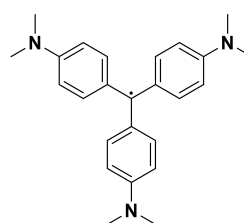
BEDT-TTF



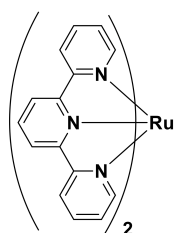
TTN



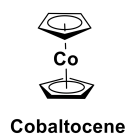
**Reduced
PPEEB**



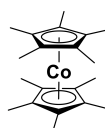
**tris[4-
(dimethylamino)phenyl]methyl
radical**



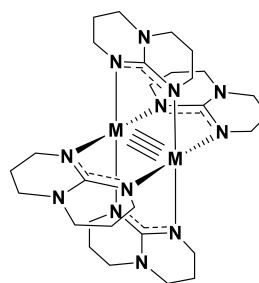
**Ru(terpy)
2⁰**



Cobaltocene



Decamethylcobaltocene



**M₂(hpp)₄
M = Cr,
Mo**

Figure 11: Examples of low ionization potential n-dopants. ^{111,114-122}

3.2.4 Organic Salts

Doping by organic salts (Figure 12) was first reported by co-evaporation of pyronin B chloride with NTCDA, and followed up shortly with a report of fullerene doping by crystal violet.^{127,128} In the latter case, mass spectroscopic and Fourier-transform infrared (FT-IR) spectroscopy data suggest that crystal violet forms leuco-crystal violet in situ. More recently, halide salts of 1,3-dimethyl-benzimidazoliums such as 2-(2-methoxyphenyl)-1,3-dimethyl-1*H*-benzimidazol-3-ium iodide (o-MeO-DMBI-I) have been used as potent n-dopants for C₆₀.^{129,130} In o-MeO-DMBI-I doped C₆₀, no iodide was found in the film by X-ray photoelectron spectroscopy (XPS) suggesting a decomposition reaction during the doping process, but this result does not exclude the incorporation of halide impurities below the detection limit of the instrument.

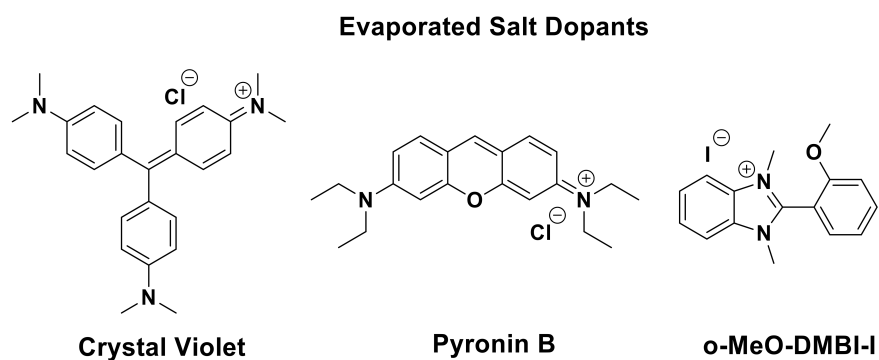


Figure 12: Examples of organic salt n-dopants.^{128,129,131}

3.2.5 Hydrides

The first report on possible hydride doping (Figure 13) was that of leuco-crystal violet processed in vacuum.¹²⁸ Similarly, pyronin B depositions may form the hydride donor leuco pyronin B.¹²⁷ More recently, 2,3-dihydro-1*H*-benzimidazoles (DMBI-H) compounds based on hydride-reduced benzimidazolium compounds were used as dopants for [6,6]-

phenyl-C₆₁-butyric acid methyl ester (PCBM) and poly[N,N'-bis(2-octyl-dodecyl)-1,4,5,8-naphthalenedicarboximide-2,6-diyl]-alt-5,5'-(2,2'-bithiophene) (P(NDI2OD-T2)).¹³²⁻¹³⁶

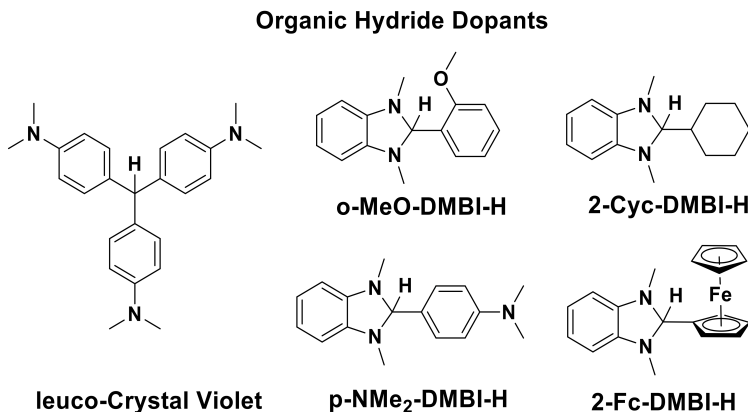


Figure 13: Examples of organic hydride donor n-dopants.^{25,128,134,137}

The mechanism of the doping reaction between DMBI-H dopants and PCBM was explored in solution, and it was determined that the doping process begins with a hydride transfer.¹³⁸ Subsequent electron transfer steps result in the formation of PCBM radical anions. This aspect of DMBI-H and related hydride dopants makes it difficult to predict their doping effect due to the unknown hydrogenation thermodynamics/kinetics with most acceptors. Additionally, the hydride donor dopants may introduce hydride-reduced impurities with unpredictable effects on charge transport.

3.2.6 Anion Doping

Studies on anion π -interactions between anhydrous halide ions and acceptors with strong positive quadrupole moments have inspired several groups to exploit these interactions for n-doping.¹³⁹ The UV-Vis spectra of fluoride doped fullerenes are consistent with the formation of fullerene radical anions.¹⁴⁰ It has been proposed that the fullerene radical anions form as a result of direct electron transfer between fluoride and fullerene. However, Weber et al. provided evidence in support of an alternative nucleophilic mechanism

by which fluoride anions generate fullerene radical anions;¹⁴¹ though, is not clear how broadly this mechanism will apply for different anion/host systems. A detailed study on the interaction of a perylenediimide acceptor (PDI) with various anions presented data that was consistent with a net thermal electron transfer between the PDI and F^- and HO^- in several aprotic solvents, but no products of electron transfer were detected for weaker Lewis basic anions.¹⁴² Similar results have been observed by Ballester et. al. between fluoride and the strong electron acceptor 1,4,5,8,9,12-Hexaazatriphenylene ($HAT(CN)_6$); though, it was proposed that the host radical anions formed by way of a charge transfer complex followed by dissociation into a solvent caged radical pair and finally loss of the fluoride radical to sacrificial reactions that prevent back electron transfer.¹⁴³ Despite the consistency of both of these studies with a net thermal ET reaction between fluoride and pi-acceptors they do not preclude the possibility that the organic radical anions form by reaction with an undetected intermediate. Nevertheless, and perhaps a more relevant detail for those interested in the electronic properties on anion-doped organic electronic devices, Li et. al. observed similar conductivities for fullerene doping by TBAX salts ($X^- = F^-, Br^-, I^-, AcO^-, HO^-$) whether or not radical anions were detected in solution for a given anion.¹⁴⁰

3.2.7 Dimers

Dimers of organic radicals (Figure 15) were first proposed as dopants in a patent.¹⁴⁷ The dimer n-dopant concept is based upon the idea that an air-sensitive dopant radical can be stabilized by dimerization through a weak C-C σ bond. Thus, radicals that are oxidized at very negative potentials can be stabilized in air at a potentially low energetic cost if the C-C bond is itself very weak.

The main advantages of this class of dopants are their air-stability and ability to dope weak electron acceptors. They are both solution and vacuum processed, and the size of the cation is tunable to enable some control over ion diffusivity, deposition rates, and

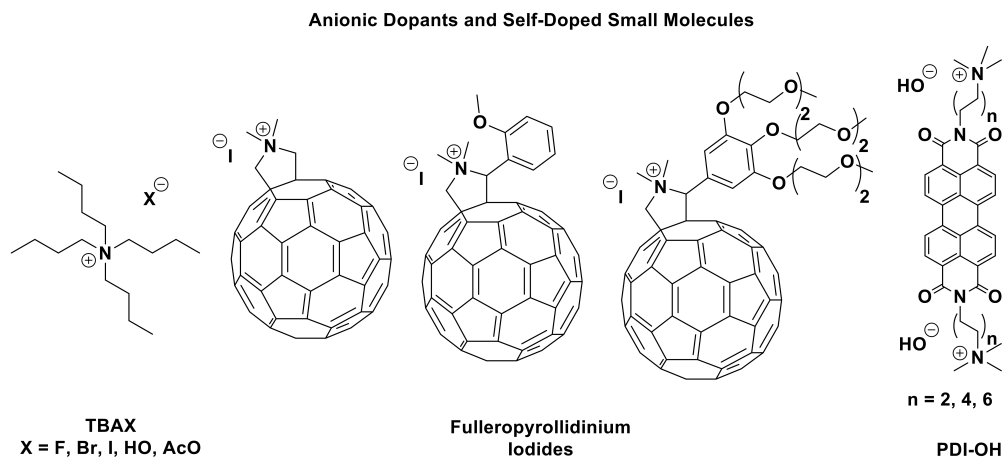


Figure 14: Examples of anionic n-dopants and self n-doped small molecules. ^{140,144-146}

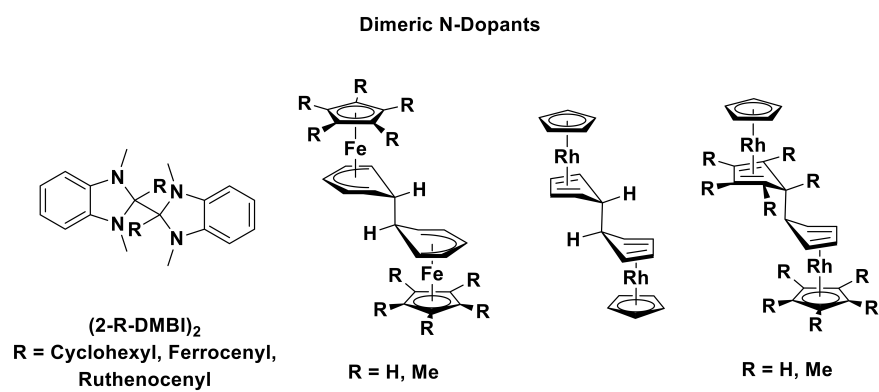


Figure 15: Examples of (DMBI)₂ and various metallocene dimer n-dopants.

morphological effects. At this time, most dimer n-dopants are synthesized by alkali metal reductions which presents a barrier to their widespread adoption.

Dimers of 19-electron metallocenes were first reported as n-dopants by the Marder and Kahn groups.^{148,149} This work was extended to dimers of organic radicals of DMBI compounds by the Bao and Marder groups.¹³⁴ A comparison of the doping effect of DMBI-H and (DMBI)₂ analogs revealed faster formation of the host anions and a more reliable doping effect between hosts with similar electron affinities for the (DMBI)₂ dopants in comparison to the DMBI-H analogs.¹³⁴ Remarkably, several organometallic and (DMBI)₂ dimers have effective ionization potentials similar to that of decamethylcobaltocene $E^{0/-} = -1.86$ V vs Fc/Fc⁺, but despite their strong reducing character, they can be handled in air as solids.^{150,151} A detailed study of the metallocene dimers by Barlow et. al. revealed that the metallocene dimers react in solution with hosts predominantly via either electron transfer first followed by cleavage of the dimer or cleavage of the dimer first followed by electron transfer, and in some cases both mechanisms may be operative.¹⁴⁸

4 Doping of Polycrystalline Organic Semiconductors

Organic doping is most commonly used in OLEDs⁴⁸ and OSCs.¹⁵² These devices consist of amorphous layers only and thus existing reviews almost exclusively summarize doping in amorphous materials (e.g.²⁸). In these materials both, the charge carrier concentration and the charge carrier mobility increases with increasing doping concentration, which leads to a superlinear increase in conductivity.²⁸

Doping crystalline or polycrystalline materials is by far more challenging to describe. A delicate interplay between crystal structure, charge carrier concentration, and trap filling leads to a complex dependency of the conductivity on the doping concentration.

Therefore, following the discussion of the chemistry of the dopant, this section focuses on the host material. Recent results on doping (poly)crystalline materials used in OFETs

are summarized and most common dopant and matrix materials are presented.

4.1 Doping of P₃HT

Doping of organic polymers stems back to the late 1970's^{65,153} and is commonly seen as the starting point of organic electronics. The ability to tune the conductivity of a polymer material from highly insulating to a level close to the conductivity of metals raised significant interest and resulted in the Noble Prize in Chemistry in 2000.¹⁵⁴

Doping of the polymeric semiconductor poly(3-hexylthiophene) (P₃HT) by the p-dopant F₄-TCNQ is one of the most intensively studied systems relevant for OFETs.^{90,155-167} In its intrinsic form regioregular P₃HT is semicrystalline, e.g. it self-organizes in a lamellar structure with two-dimensional sheets with strong $\pi - \pi$ interactions surrounded by amorphous regions of insulating side chains.¹⁶⁸ This semicrystalline phase leads to efficient charge transport and relatively high charge carrier mobilities of intrinsic P₃HT.

Doped P₃HT films are usually deposited by spin-coating from a mixture of P₃HT and F₄-TCNQ. Charge transfer between P₃HT and F₄-TCNQ is readily observed in solution^{156,160} even before film formation. This charge transfer is visible as novel sub-bandgap signals in the absorption spectra, which can be assigned to F₄-TCNQ anions, P₃HT bipolarons, and P₃HT aggregates.¹⁵⁶ The charge transfer in solution is efficient. Pingel et al.¹⁶⁰ and Wang et al.¹⁵⁶ estimate that about 60% of all F₄-TCNQ molecules are ionized in solution. Furthermore, the observation of a strong F₄-TCNQ anion signal indicates that an integer type charge transfer is more likely¹⁶⁰ than the formation of an hybrid state involving partial charge transfer.

The interaction between F₄-TCNQ and P₃HT was found to be highly localized. From absorption spectra taken at different doping concentrations Wang et al. were able to show that one F₄-TCNQ anion interacts with approximately 4.9 P₃HT monomers.¹⁵⁶ This is in agreement with an earlier study showing that charge transfer complexes formed in

doped P₃HT are localized to about four thiophene units.^{159,163}

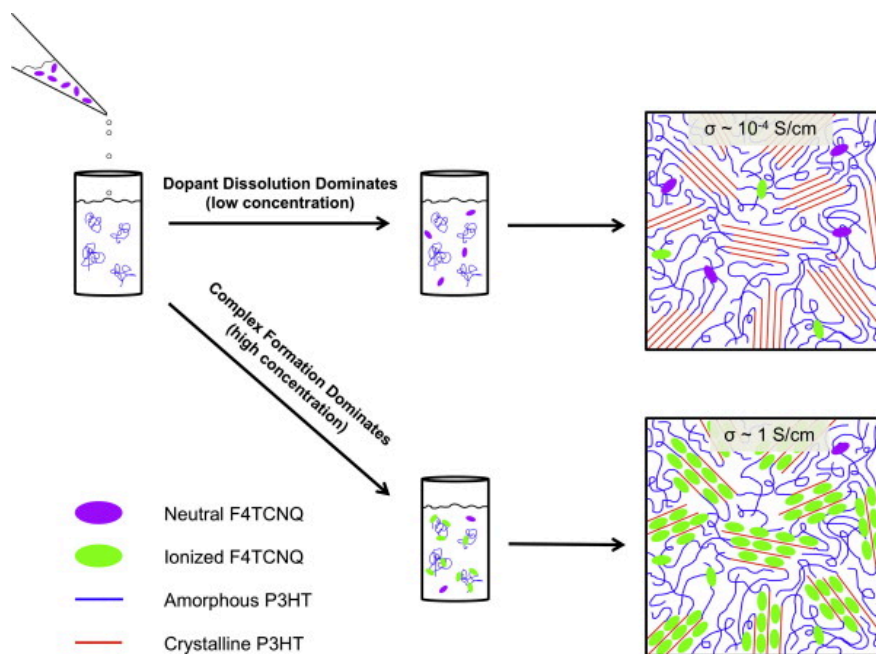


Figure 16: Weak and strong doping regime in P₃HT. At low doping concentrations F₄-TCNQ anions are mainly incorporated into amorphous regions of the film. At higher concentrations, a mixed crystalline phase consisting of P₃HT and F₄-TCNQ is formed. Reprinted with permission from⁹⁰. Copyright 2013 Elsevier.

Thin doped P₃HT films can be formed by spin coating. The structure of these doped P₃HT films depends on the concentration of F₄-TCNQ (see Figure 16). At low doping concentrations and increased temperature during spin coating the structure of the doped films resembles the structure of undoped P₃HT.⁹⁰ Duong et al. argue that in this so called weak doping regime the dopant ions are incorporated into the amorphous region of the film. Charge transfer to P₃HT is weak but still visible in highly sensitive photothermal deflection spectroscopy spectra.⁹⁰

At increased concentrations F₄-TCNQ anions are incorporated into the crystalline region of the film leading to a distinct change in film structure. In particular, the lamellar spacing increases.⁹⁰ Mixed crystals coexist with pure P₃HT phases in the film and contribute to a strong increase in conductivity.¹⁵⁸ Hence, this phase was denoted as the strong doping phase. If the lamellar structure of P₃HT is suppressed, as e.g. in

regio-random P₃HT, doping was found to be much less efficient.¹⁵⁷

At even increased doping ratios the crystalline phase of P₃HT becomes saturated of F₄-TCNQ anions and pure F₄-TCNQ crystals form. The onset of the saturation corresponds well with the observation that one F₄-TCNQ anion interacts with approx. four P₃HT monomers.⁹⁰

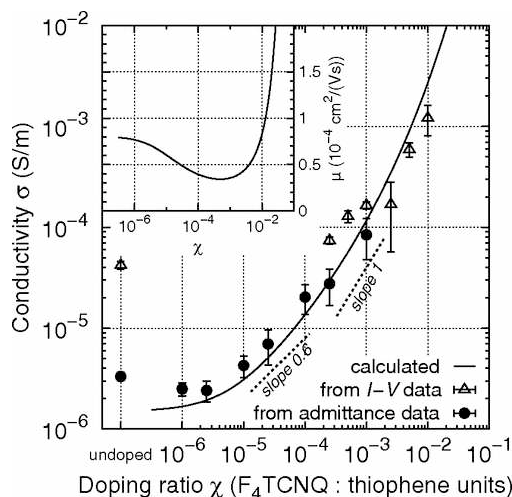


Figure 17: Conductivity and charge carrier mobility in doped P₃HT films. At low doping concentrations the mobility drops, which is explained by a broadening of the density of transport states. This trend is reversed at large doping concentrations, i.e. the charge carrier mobility increases. Reprinted with permission from¹⁶⁰. Copyright 2013 American Physical Society.

The changes in structure are reflected in the charge transport properties intensively studied by Pingel and Neher^{160,161} (see Figure 17). Although charge transfer from F₄-TCNQ to P₃HT monomers is highly efficient, only about 5% of all F₄-TCNQ anions contribute a free hole to P₃HT.¹⁶¹ This effect, commonly observed in other materials as well and discussed in section 2, is explained by a strong coulomb interaction between the F₄-TCNQ anion and positively charged P₃HT monomer leading to a large activation energy of doping and hence a low probability of activation of dopants.

At intermediate doping concentrations Pingel and Neher observe a decrease in charge carrier mobility.¹⁶¹ Furthermore, by Kelvin potential measurements they observe a broadening of the density of states in P₃HT upon doping, which is explained by the

analytical model of Arkhipov et al.¹⁶⁹ They argue that randomly distributed dopant anions coulombically interact with free charges, which increases the energetic disorder in the film and thus broadens the density of transport states. In particular low lying states act as trapping sites, effectively decreasing the charge mobility.

At larger doping concentrations, the trend is reversed and an increase in mobility is observed.¹⁶⁰ This effect is rationalized by a coupling of the coulomb potential of dopant ions leading to a lower effective activation energy as proposed by Mityashin et al.³⁶ and a substantial filling of low lying states in the density of states by donated holes. Thus, charge is transported at an increased density of states, leading to higher hopping rates and increased mobility.

4.2 Doping of Pentacene

If P₃HT is a model compound for a doped polymer, pentacene is the archetype of a high mobility small molecular organic semiconductor. For long, it set the record in charge carrier mobility for vacuum processed OFETs. It is in the range of the mobility of amorphous silicon ($1 \text{ cm}^2/\text{Vs}$),^{170,171} which triggered strong interest into this material early on in the development of OFETs.

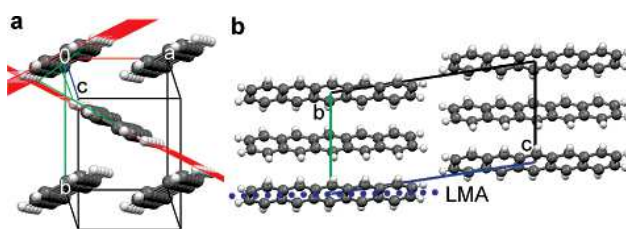


Figure 18: Structure of the thin film polyomorph of pentacene. Pentacene crystallizes in a layered structure with the long axis of the molecule perpendicular to the substrate. Large orbital overlap in the ab plane due to the herringbone structure leads to efficient charge transport. Reprinted from¹⁷². Copyright 2007 American Chemical Society.

The high mobility of pentacene can be explained by a crystal structure favorable for charge transfer between adjacent pentacene molecules. Pentacene crystallizes in the triclinic unit cell shown in Fig. 18. In thin films, the long axis of the molecule is

aligned approximately perpendicular to the substrate. Within the ab plane, the molecules adopt a herringbone structure, which maximizes the orbital overlap along the lamella of the herringbone structure, which results in a high mobility in the plane parallel to the substrate.

Several different polymorphs of this thin film structure are reported, but they are assumed to share the herringbone arrangement.¹⁷³ They differ slightly in the tilt angle of the pentacene crystal and therefore in the layer periodicity $d(001)$, i.e. in the distance between the ab planes.

4.2.1 P-Doping of Pentacene by Iodine

Doping of pentacene has been studied since the early 90's.¹⁷⁴⁻¹⁸⁰ In these early reports, p-doping of pentacene was accomplished by exposing the pentacene films or single crystals to iodine vapor.¹⁷⁴ Iodine is incorporated into the pentacene crystal and electrons are transferred from pentacene to iodine, leaving behind a hole on the pentacene crystal. This charge transfer was proven e.g. by UV-Vis absorption, which shows the appearance of new absorption bands in doped films assigned to the formation of ionized iodine and pentacene,¹⁸¹ and Raman spectroscopy, which shows signatures of I_3^- and I_5^- species in the film.^{181,182}

Iodine doping increases the hole density in the film. It has been estimated to be rather high - up to 10^{21} cm^{-3} at high doping ratios.¹⁷⁷ Furthermore, as the charge carrier mobility is not significantly reduced upon doping, a high conductivity of up to 150 S/cm is reached.¹⁷⁷ Even a transition to metallic behavior was proposed,¹⁷⁷ which, however, was questioned later on.¹⁸⁰

The high mobility and conductivity in these doped films implies that the crystal structure remains largely undisturbed even at higher doping concentrations. The details of the structure of iodine doped pentacene was studied by X-Ray diffraction (XRD),^{178,183} and Raman and IR spectroscopy.¹⁸² XRD data shows that the distance between the ab

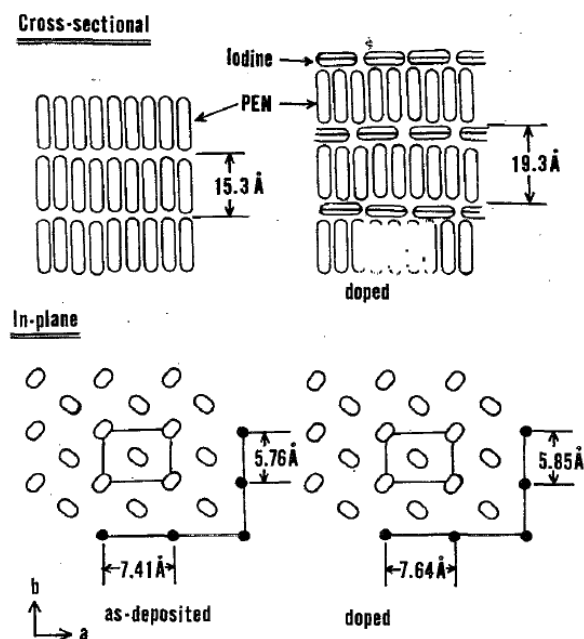


Figure 19: Doping pentacene by Iodine. Iodine accumulates between the layers of pentacene to form an intercalated structure. The high in plane mobility is preserved. Reproduced from ¹⁷⁸, with the permission of AIP Publishing.

planes of pentacene molecules is increased, whereas the molecular structure in the *ab* plane remains almost unchanged. For example, Minakata found an increase from 15.3 Å of the undoped film to 19.3 Å of the doped film.¹⁷⁸ This "swelling" of the film has been interpreted as an intercalation of iodide ions in between the layers of pentacene (see Figure 19). Considering that pentacene shows its highest mobility in the *ab* plane, which remains almost undisturbed, it is not surprising that the in-plane mobility remains high. However, as observed by Minakata et al. as well, doped pentacene films show a huge anisotropy between the out of plane and in plane conductivity, which is as high as 8 orders of magnitude.¹⁷⁴

This model of intercalation of iodine between layers of pentacene was confirmed by a more detailed study of Ito et al.¹⁸³ They found two transitions of pentacene upon doping. At intermediate doping concentrations they observed a transition from the triclinic unit cell of undoped pentacene to an orthorhombic one. In agreement with Minakata, the

lattice constants in the ab plane are nearly unaltered, but the lattice parameter c is increased from 14.5 Å to 19.28 Å. This phase was identified as the highly conductive phase reported by Minakata. However, they found an additional phase at higher doping concentrations, where the in-plane lattice constants are increased, which they explain by an incorporation of iodine into the pentacene layer. Raman and IR spectroscopy confirm the presence of the additional structure at high doping concentrations.^{181,182} Due to the increase of the ab lattice constants, the in plane mobility and conductivity of this phase decreases.¹⁸¹

Overall, iodine doping of pentacene is highly effective. As iodine is mainly intercalated between the planes of pentacene, leaving the in-plane herringbone structure intact, doping does not deteriorate charge carrier mobility, which results in high conductivities and effective charge transport. However, iodine doping comes at severe disadvantages, which render it impractical for real devices. First of all, exposing the pentacene film to a vapor of the dopant leads to inhomogeneous doping profiles, which are difficult to control.¹⁸¹ Furthermore, the doping effect is not stable, i.e. the incorporated iodine is desorbed from the layer in ambient conditions, which leads to a drop in conductivity over time.¹⁷⁷

4.2.2 P-Doping of Pentacene by Molecular Dopants

Although highly efficient, iodine tends to diffuse in pentacene, rendering the doping effect unstable and precluding use of iodine in organic devices. This shortcoming has fueled research in organic dopants, which are more bulky and do not diffuse in the organic host layers.²⁸ However, the increased stability comes at a high price for polycrystalline materials. Due to the larger size of the dopant, it is expected that the crystal structure of the matrix material is disturbed leading to a reduction in charge carrier mobility.

The most commonly used molecular dopant is F₄-TCNQ.⁷ It is heavily used to increase the conductivity of hole transport layers of OLEDs⁴⁸ and organic solar cells.²⁸ Furthermore, it is known to dope polycrystalline pentacene as well.^{34,184}

Ha and Kahn studied the microscopic structure of F₄-TCNQ doped pentacene films by scanning tunneling microscopy (STM).³⁸ They found that F₄-TCNQ molecules diffuse into vacancies of the pentacene lattice and were able to show that the dopant does not disturb the crystal lattice of pentacene.

However, Ha et al. studied low doping concentrations only. At higher concentrations it was shown that the crystallinity of pentacene is indeed reduced.¹⁸⁵ For concentrations above 1wt.% the average crystallite size decreases (see Figure 20), which results in a continuous decrease in charge carrier mobility. This decrease is even stronger for larger dopants such as F₆-TCNNQ.

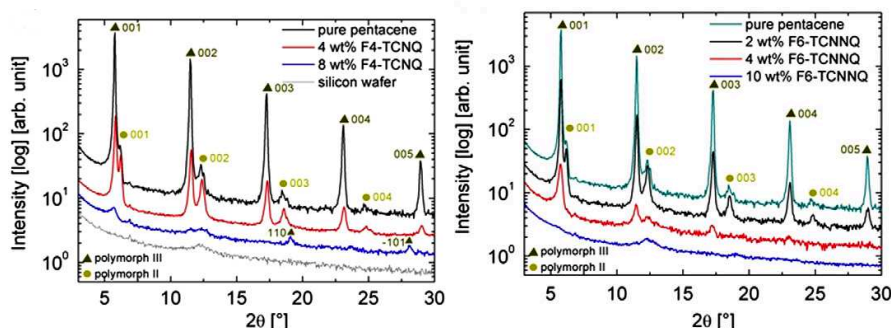


Figure 20: Doping pentacene by F₄-TCNQ and F₆-TCNNQ. XRD shows weaker reflections at higher doping concentrations, corresponding to smaller crystallite size. Reprinted with permission from ¹⁸⁵. Copyright 2012 Elsevier.

This reduction in crystallinity leads to a drop in mobility. Consequently, the conductivity of a doped pentacene layer increases at first due to the increase in charge carrier concentration. However, it reaches a maximum of approx. 0.1 S/cm at 6 wt.% (F₄-TCNQ) or 2 wt% (F₆-TCNNQ), before it drops due to the decrease in charge carrier mobility induced by doping.

This increase in disorder at higher doping concentrations was as well observed for another p-dopant, C₆₀F₃₆,⁵⁰ by Pahner et al.⁶¹ They report on a broadening of the density of occupied states at higher doping concentrations, which can be understood as a direct result of increased disorder. However, in accordance to Ha et al.,³⁸ no broadening was observed at low doping concentrations.

Most interestingly, Pahner et al. were as well able to deduce the concentration of free charge carriers in the doped film from the capacitive response of a Schottky diode. They observed that the density of free charge carriers is thermally activated and found an activation energy of 54 meV for $C_{60}F_{36}$ and 19 meV for F_6 -TCNNQ. This activation energy of $C_{60}F_{36}$ in pentacene was confirmed qualitatively by UPS measurements.⁴⁷ As discussed in Section 2 this activation energy can be explained by Coulomb interactions between transferred charges³⁶ or activation from the bonding state of dopant-pentacene hybrid.³⁹

4.2.3 P-Doping of Pentacene by Transition Metal Oxides

Transition metal oxides are well known dopants used in OLEDs and organic solar cells.⁷⁸ These materials exhibit a very large electron affinity of around 6.7 eV (MoO_3 and V_2O_5) or 6.5 eV (WO_3), which makes them capable of doping wide gap materials.

Several authors report on doping pentacene by MoO_3 .^{23,186–188} Charge transfer from dopant to pentacene was proven by UPS²³ and XPS.¹⁸⁷

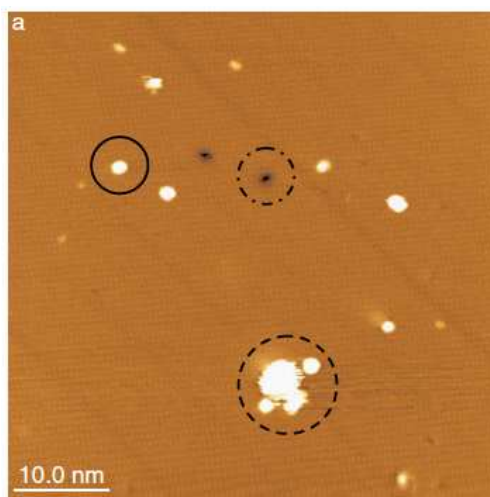


Figure 21: Doping pentacene with MoO_3 . Clusters of MoO_3 are seen in STM images of a crystalline pentacene surface. Reprinted with permission from¹⁸⁶. Copyright 2010 American Physical Society.

As for F_4 -TCNQ, Ha et al. studied doping pentacene by MoO_3 using STM.¹⁸⁶ They deposited a low amount of MoO_3 on top of a crystalline pentacene surface. In contrast to

doping with F₄-TCNQ, MoO₃ does not diffuse towards vacancies in the pentacene lattice and smaller and larger clusters of MoO₃ are formed on the pentacene surface (see Figure 21). This clustering is in accordance with the observations of Berkowitz et al., who found that MoO₃ forms clusters already in the gas phase when sublimed.¹⁸⁹

Ha et al. observe that a layer of pentacene grown on the clusters of dopants exhibits the regular crystal structure, which indicates that doping by MoO₃ does not degrade the structural order of the film, at least not at low concentrations. For higher concentrations (above 0.5 mol%) a reduction in order²³ and crystallite size¹⁸⁷ is observed.

4.2.4 N-doping of Pentacene

N-doping of pentacene is by far more challenging than p-doping. Due to the low electron affinity of pentacene (approx. 2.72 eV¹²⁴), the n-dopant has to show a very low ionization potential to allow for an electron transfer, which renders most potential candidates unstable in air. However, as will be discussed below, air-stable and effective n-dopants for pentacene were found as well.

In first studies, n-doping of pentacene by alkali metals such as sodium, potassium, and rubidium,¹⁷⁹ lithium¹⁹⁰ or cesium¹⁹¹ was studied. All metals show a doping effect in pentacene. A conductivity of 10⁻⁵ S/cm is found for Na doping,¹⁷⁹ 6 × 10⁻³ S/cm for lithium,¹⁹⁰ 100 S/cm for potassium,¹⁹² 47 S/cm for rubidium¹⁹¹ and 74 S/cm for cesium doping.¹⁹¹ This high conductivity and the observation that the conductivity increases with decreasing temperature is sometimes even interpreted as a metallic transport in these mixed layers.¹⁹¹

For many alkali metal dopants an intercalated structure as for iodine doping of pentacene was found as well. For example Hansson et al. report on a molecular dynamics (MD) and density-functional theory (DFT) study of the structure of potassium doped pentacene.¹⁹³ At low concentration, potassium intercalates between the pentacene layers, leaving the herringbone structure almost unchanged. At higher concentrations, however,

the herringbone structure vanishes due to an incorporation of ions into the ab plane. This leads to a strong $\pi - \pi$ -interaction along the c-axis, which should increase the out-of-plane mobility.

However, despite the high conductivity observed in these films, doping by alkali metals has significant drawbacks. In particular, it is challenging to control doping ratio (the pentacene films are often exposed to a vapor of the alkali metal) and the doped films often show a low stability. Finally, the high concentrations of alkali metals sometimes used to prepare films with high conductivity (up to a 1:1 ratio) show that rather than having a doped film of pentacene, a mixed hybrid film of pentacene with alkali metal is studied, which makes the interpretation of results challenging.

These shortcomings can be avoided by molecular dopants. Chan et al. reported on the doping effect of the strong electron donor decamethylcobaltocene CoCp^*_2 in pentacene in 2008.¹²⁴ Although the ionization energy of CoCp^*_2 is larger than the electron affinity of pentacene, the generation of free electrons on the pentacene matrix was proven by UPS. By capacitance-voltage measurements they were able to estimate the concentration of free electrons in the film and showed that the doping efficiency, i.e. the ratio of free electrons to dopant molecules is only about 1%.

Still, CoCp^*_2 is highly volatile and doping is achieved by exposing the sample to a background pressure of the dopant during growth of the film. Although it was shown that the dopant is indeed incorporated into the pentacene film,¹²⁴ a dopant with a lower vapor pressure is desirable as it would allow for a better control in doping concentration.

A larger dopant, bis(4,4',4''-tri-tert-butyl-2,2':6',2''-terpyridine)ruthenium ($[\text{Ru}(\text{t-but-terpy})_2]^0$), was proposed by Harada et al.³⁴ Doping was studied by UPS and thermovoltage measurements. It could be shown that n-doping shifts the Fermi level of the doped layers towards the LUMO states, as would be expected from classical semiconductor statistics. Harada et al. combined n- and p-doping by $\text{F}_4\text{-TCNQ}$ to process a p-i-n homojunction and were able to show that the built-in voltage of the junction (1.65 V) is

related to the open circuit voltage if the junction is operated as solar cell.

$[\text{Ru}(\text{t-but-terpy})_2]^0$, as most n-dopants, is air sensitive and has to be handled under protective environments. In 2012 Guo et al. proposed a novel concept for air-stable n-dopants, which are strong enough to dope pentacene.^{148,149,194} Based on their results on doping pentacene by decamethylcobaltocene,¹²⁴ they synthesized organometallic dimers of rhodocene ($[\text{RhCp}_2]_2$) and similar dimers, e.g. ruthenium(pentamethylcyclopentadienyl)(1,3,5-triethylbenzene) $[\text{Cp}^*\text{Ru}(\text{TEB})]_2$ (see Figure 15). In contrast to the monomers, the dimers are relatively air-stable, have a lower vapor pressure, and can be co-evaporated with the matrix material, allowing for a precise control of the doping concentration.

These dopants were shown to be capable of doping a wide range of matrix materials, including pentacene and TIPS-pentacene. The current density was shown to increase greatly in doped layers and the Fermi level shifts toward the LUMO.¹⁴⁹ Two mechanisms are discussed to explain doping. In the first mechanism, the dimer is cleaved first, which alone is an endergonic or unfavorable process. However, this process is followed by a rapid charge transfer to pentacene, stabilizing the result. In the second mechanism, it is assumed that the electron transfer from dopant to pentacene happens first, which is followed by a cleavage of the dimer. Depending on the choice of dopant and matrix materials, both mechanisms can be in effect.¹⁹⁴

In some of these doping studies it is shown that pentacene possesses a significant density of electron traps³⁴ and that filling these traps by doping can inactivate these traps electronically.¹⁴⁹ This effect was used to prepare films of 6,13-bis(triisopropyl-silylethynyl) (TIPS)-pentacene with a very high electron mobility. Doping these films by o-MeO-DMBI, which is not strong enough to supply free electrons to the LUMO of TIPS-pentacene, leads to a filling of electron traps located in the HOMO/LUMO gap. It was shown that filling these traps leads to highly efficient electron transport and a mobility of $3.2 \text{ cm}^2/\text{Vs}$ on average.¹³³ Furthermore, the crystal structure of the doped films is not disturbed by the dopant. Most probably a phase segregation is taking place, which is in accordance to

the relatively large doping concentrations needed (50 mol%).

4.3 Doping of Fullerene C₆₀

Whereas pentacene is the archetype of an organic p-type semiconductor, C₆₀ and its soluble counterpart PCBM is heavily used as n-type channel material in OFETs.

N-doping of C₆₀ by alkali metals was found to be very efficient in 1991 by Haddon et al.¹⁹⁵ Furthermore, superconductivity was observed for potassium doped films.^{196,197}

Haddon and Hebard exposed films of C₆₀ to vapors of Cs, Rb, K, Na, or Li, which lead to a pronounced increase in conductivity.¹⁹⁵ A conductivity of 500 S/cm was reached for the most efficient dopant, potassium. The authors explain this observation by a doping effect and argue that the crystal structure of C₆₀ leaves enough free space for alkali metal ions to incorporate into the structure. However, at even higher concentrations the conductivity started to decrease. In particular for the largest ion, Cs, the film darkened and became rough, which was understood as a disruption of the crystal lattice due to difficulties of incorporating too many Cs ions.

Although n-doping by alkali metals is highly efficient, the same shortcomings as for pentacene apply. In particular the low stability of these films precludes the use of these dopants in organic devices. However, many larger molecular dopants were reported in literature and many compounds were proven to be effective. For n-doping, rhodamine B,¹²⁷ Ca,¹⁹⁸ ruthenium(pentamethylcyclopentadienyl)(1,3,5-trimethylbenzene) dimer ([RuCp*(mes)]₂),^{51,52} N-DMBI,²⁵ 2-(1,3-dimethyl-1H-benzoimidazol-3-ium-2-yl) phenolate hydrate (DMBI-POH),¹³⁰ acridine orange base (AOB),^{131,199} leuco-crystal violet,¹²⁸ and W₂(hpp)₄^{120–122} were proposed.

Although a small reduction in charge carrier mobility with increasing n-doping concentration was found,¹⁹⁹ in most reports doping did not degrade the charge transport properties or the crystallinity of the samples.

Instead, even an increase in mobility with the doping concentration is found, which

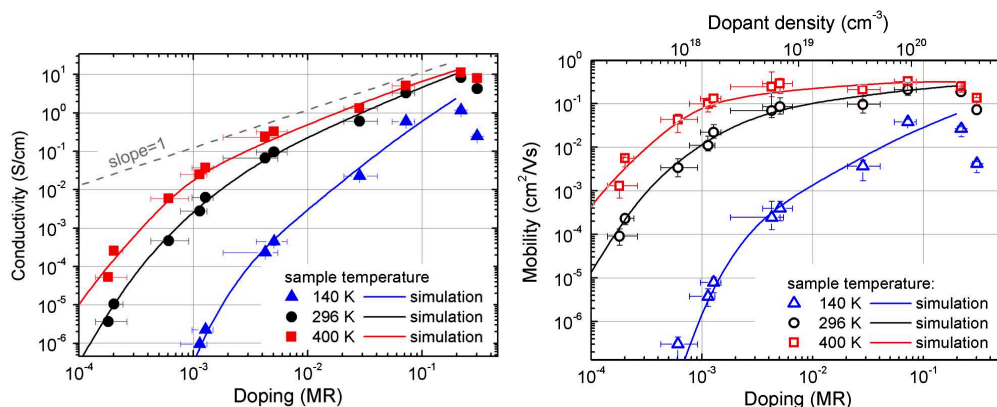


Figure 22: Conductivity and charge carrier mobility of films of C_{60} doped by the n-dopant $[RuCp^*(mes)]_2$. A superlinear increase in conductivity and increase in mobility is found for low doping concentrations, which is due to a filling of traps states by dopants. Reprinted with permission from⁵². Copyright 2012 American Physical Society.

is a remarkable difference to p-type doping of pentacene^{51,52} (see Figure 22). This effect was explained by filling of trap states, which hinder electron transport in intrinsic layers. Filling these traps is assumed to electrically passivate traps, which leads to more effective charge transport⁵² and even to an increase in air stability of the doped films.¹³² Only at very high doping concentrations (molar ratio of 20%), a decrease in conductivity and mobility is found (see Figure 22), indicating a reduction in order in these films.

Most of the n-dopants reported in literature degrade rapidly upon exposure to air. However, recently it was shown that doping C_{60} by the air-sensitive n-dopant $W_2(hpp)_4$ leads to a passivation of the n-dopant, which renders the film at least partially stable in air.¹²¹ Although the conductivity of the doped film drops in air, it can be recovered by thermal annealing in vacuum. This effect was explained by electron transfer and the higher stability of the $W_2(hpp)_4$ cation.¹²⁰ It furthermore depends on the position of the LUMO of the host material. For hosts with an even larger electron affinity compared to C_{60} , e.g. N,N-Bis(fluoren-2-yl)-naphthalenetetracarboxylic diimide (Bis-Hfl-NTCDI)^{11,120} the effect is increased and yields organic films doped by an air sensitive n-dopant, which are nevertheless stable in ambient conditions.¹²⁰

P-doping is much more challenging given the low HOMO edge of approx. 6.4 eV. So

far only doping using the transition metal oxide MoO_3 has been reported.^{198,200–202} Combining p and n-type doping, Shinmura was able to prepare a C_{60} based homojunction²⁰² similar to the pentacene homojunction of Harada et al.³⁴ As for n-doping, the crystal structure of C_{60} seems not to be heavily disturbed by the dopant. For example, it was shown that doping C_{60} by MoO_3 only slightly reduced the grain size of the polycrystalline film from 9 nm to 7 nm²⁰¹ and that the dopant is evenly distributed in the film.

4.4 Doping Other Organic Semiconductors with High Charge Carrier Mobility

Until 2007 pentacene was the unbeaten record holder and exhibited the highest hole mobility (in the range of 1-3 cm^2/Vs). In the search for compounds with a higher air-stability the molecule dinaphtho[2,3-b:2',3'-f]thieno[3,2-b]thiophene (DNTT) was developed,^{13,203} which turned out to be more stable and shows a slightly increased mobility ($\mu \approx 3 \text{ cm}^2/\text{Vs}$). It was found that the mobility can be improved further, if alkane spacer chains are attached to both sides of the DNTT core (Cn-DNTT), yielding a mobility as high as 8.5 cm^2/Vs .^{204–207} Recently, an off-center spin coating technique has been presented that yields thin crystalline films of the compound 2,7-dioctyl[1]benzothieno[2,3-b][1]benzothiophene (C8-BTBT) with a maximum mobility of 43 cm^2/Vs , which is one of the highest values reported so far for organic semiconductors deposited on large areas.²⁰⁸

Although the use of well known electron acceptors such as $\text{F}_6\text{-TCNNQ}$ or $\text{F}_4\text{-TCNQ}$ was reported for some of these materials, so far the dopant is mainly deposited as a pure layer on top of the organic semiconductor to improve injection at the source and drain contacts,^{13,206,207} and doping in mixed layers is seldomly reported.

C8-BTBT or the related semiconductor 2,7-didecyl(C_{10})-BTBT is an exception. Mendez et al. studied the doping of C_{10} -BTBT by $\text{F}_4\text{-TCNQ}$ and found a unit cell consisting of dimers of the conjugated core of C_{10} -BTBT and $\text{F}_4\text{-TCNQ}$ (see Figure 23). An increase in conductivity was found at low doping concentrations, however, it remained unclear if

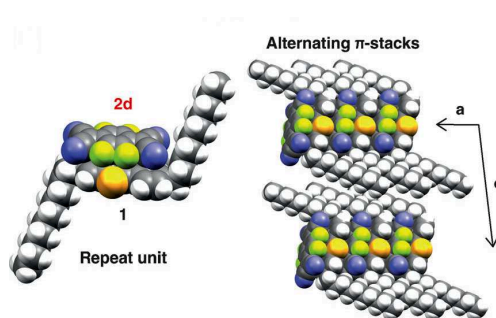


Figure 23: Doping C10-BTBT by F₄-TCNQ. Dimers of C10-BTBT and F₄-TCNQ are formed. Reprinted with permission from⁴¹. Copyright 2013 John Wiley and Sons.

F₄-TCNQ is actually providing free holes to C10-BTBT or if the increase in conductivity is due to filling of traps. Considering the results reported by Soeda et al.,²⁰⁹ who studied doping of C8-BTBT by F₄-TCNQ in organic transistors, trap filling is more probable. Soeda et al. find no increase in the off-state current of OFETs upon doping, but a reduction in threshold voltage, similar to the results of Olthoff et al. for n-doping C₆₀⁵¹ discussed above (see Section 5.2.2 for a more detailed discussion).

A similar packing motif was found by Cochran et al. for the polymer poly(2,5-bis(3-tetradecylthiophen-2-yl)thieno-[3,2-b]thiophene) (PBTTC-C14) doped by F₄-TCNQ.⁹¹ The authors found that F₄-TCNQ stacks cofacially with the conjugated backbone of the polymer forming an intercalated crystalline microphase leading to high charge transfer probabilities. Most interestingly, their data suggest that instead of a random arrangement of the dopant in the crystalline region, some polymer chains are highly decorated with F₄-TCNQ, whereas other chains are not. Thus, the distribution of dopants is not entirely random as assumed in most transport models, but correlated.

Most organic dopants disrupt the structure of the matrix material leading to a decrease in charge carrier mobility (see as well Figure 20). To minimize this disruption, Gregg et al. proposed a dopant/matrix system where the dopant is commensurate with the lattice.^{37,57,115,210} They doped the liquid crystalline semiconductor PPEEB by a derivative of the same molecule with a covalently bound positive countercharge added to the end of the alkyl chain. They observed a superlinear increase in conductivity,²¹⁰ which was

explained by a decrease in activation energy of the doped charge carriers.¹¹⁵

5 Doped Organic Transistors

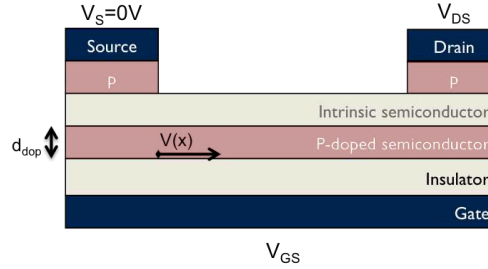


Figure 24: Setup of a p-doped organic transistor. In some reports, only a thin layer at the interface to the oxide is doped (thickness d_{dop}). To ensure efficient injection of holes, the source and drain regions can be p-doped as well.

The general setup of a p-type OFETs is shown in Figure 24. From bottom to top, the transistor consists of a metallic gate electrode covered by an insulating layer and a thin layer of a p-type organic semiconductor. The organic semiconductor is furthermore contacted by two electrodes - the source and drain electrodes. To operate the OFET a voltage is applied between the source and the drain electrode (V_{DS}) and the current leaving the transistor at the drain I_D is measured. The drain current is modulated by the voltage applied between the source and the gate electrode V_{GS} . If, for example, a negative voltage is applied to the gate with respect to the source electrode, holes are electrostatically accumulated at the interface between the organic semiconductor and the gate insulator. As the conductivity of the organic semiconductor increases with the charge carrier density, the drain current increases as well, i.e. the transistor switches on. If, however, a positive voltage is applied to the gate, all free holes are repelled from the gate insulator. The semiconductor is fully depleted of charge carriers, the local conductivity drops to zero, and the transistor switches off.

Doping can improve OFETs in various ways. It can be used to increase the saturation current and transconductance of OFETs, to fine-control transistor behavior, and to make

transistors more reliable.

Contact doping as reviewed in Section 5.1 is the most common use of doping in transistors,^{13–17,19,20,211} and has been recently identified to be essential for further improving the switching frequency of organic transistors.²¹¹

Besides contact doping, doping of the channel of the transistors^{21–24,166,209} leads to a precise control of the threshold voltage and to an increase in stability.^{26,132} Details and limits of channel doping will be discussed in Section 5.2.

Last but not least, doping adds a further dimension in the control of the electric properties of organic semiconductors. In particular the control of the Fermi level position (see e.g. Figure 6) allows to control the majority charge carrier species,^{133,201} to realize inversion OFETs based on minority charge carriers,^{21,212} or to control ambipolar behavior of transistors²¹³ (Section 5.2.3).

5.1 Contact Doping

Efficient charge injection is essential for high-performance of OFETs since it is the starting point of the electrical operation. Even if organic semiconductors with extremely high charge carrier mobility are used, the performance of organic transistors will be severely limited by inefficient injection.

In most OFET architectures, metal electrodes are utilized to establish contact with the organic semiconductor layer and to inject charges into (or extract charges from) it. Therefore, charge injection is mostly dominated by the interfacial properties between the metal electrodes and organic semiconductors. It is of crucial importance to not only develop a physical model of charge injection but to understand all physical phenomena occurring at a metal/organic interface as well. Hence, enormous efforts have been devoted to describe injection into organic semiconductors and to characterize the electronic structure and the energy alignment at the interface.^{12,214–222}

The energy level structure of the junction formed between metal and organic semi-

conductor is schematically sketched in Figure 25. When two heterogeneous materials are brought into contact, a thermodynamic equilibrium is established via charge carrier transfer across the interface. This process commonly gives rise to a space charge region at the metal/organic interface, leading to a bending of the energy levels depending on the relative energy alignment of metal work function (ϕ_m) and Fermi level (E_F) of the organic semiconductor. Interface dipoles (Δ) induced by the interaction between the organic semiconductor and the metal, will modify the metal work function, which additionally alters the electronic structure of the interface.²²³

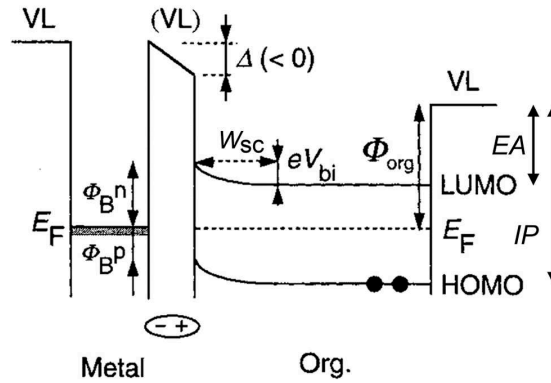


Figure 25: Interfacial energy diagram at the metal/organic junction. Adapted with permission from²¹⁴. Copyright 1999 John Wiley and Sons.

The contact resistance impeding efficient charge injection originates from potential barriers formed at the interface. The energy barriers at the interface for hole (ϕ_B^p) and electron (ϕ_B^n) injection shown in Figure 25 are given by

$$\phi_B^p = IP - \phi_m - \Delta \quad (6)$$

$$\phi_B^n = \phi_m - EA + \Delta \quad (7)$$

A straightforward strategy to improve charge injection is to eliminate the energy barrier at the metal/organic interface so that the contact resistance is minimized and a so-called Ohmic contact is formed. According to Equations 6 and 7, the charge injection barrier can be minimized by matching the metal work function to the charge transport

level i.e. HOMO and LUMO of the organic semiconductors.

However, most metals that would enable efficient electron injection are usually highly reactive rendering the transistor unstable in ambient conditions. Therefore, a wide range of approaches to engineer contacts and/or tune the metal work function were introduced. For example, self-assembled monolayers (SAMs)^{222,224–226} or other injection layers were used to minimize injection barriers.^{227–230} Among these techniques, contact doping was studied extensively and was proven to be a promising approach to realize quasi-Ohmic contacts.

In this section, we summarize the fundamental theory of the contact resistance and contact doping, and review recent studies on contact doping used in wide variety of transistors.

5.1.1 Contact Resistance

The total resistance R_{tot} of OFETs is usually modeled as a series connection of the contact resistance R_c and the channel resistance R_{ch} as illustrated in Figure 26 for an inverse staggered configuration,

$$R_{tot} = R_{ch} + R_c \quad (8)$$

When the channel length L of OFETs is reduced, R_{ch} decreases as well, whereas the contact resistance remains unchanged as it is inherently governed by interfacial properties. Thus, the overall electrical performance of short channel devices is limited by the contact properties.

R_c consists of two contributions, the interface resistance R_{int} determined by the energy barrier at the metal/organic interface and the bulk resistance R_{bulk} , i.e. the resistance of the organic material itself.¹⁹ In the so-called coplanar architecture, where the charge accumulation channel and source/drain electrodes lie in the same plane, i.e. at the semiconductor/dielectric interface, R_{bulk} can be neglected since charges are directly injected into the channel. However, it has been shown that charge injection in

coplanar devices is generally poorer than in the staggered structure. This difference was rationalized by the argument that in the staggered configuration injection can occur through the entire contact area defined by the overlap of the source/drain area with the gate electrode.^{231–234}

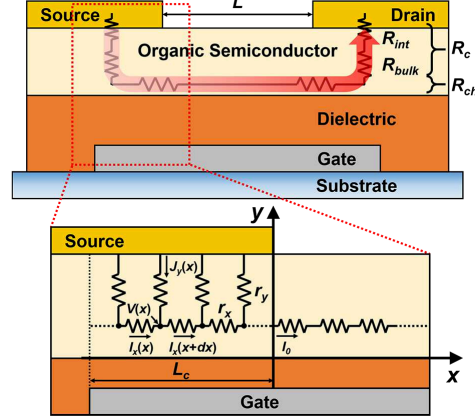


Figure 26: Simplified modeling of the total resistance in an inverse staggered OFET (top-contact bottom-gate structure) (top) and equivalent circuit of the current crowding model for extracting the contact resistance (bottom).

In 1969, Murrmann and Widmann proposed an analytical model named "current crowding model" to estimate the contact resistance between metal and semiconductor.^{235,236} As illustrated in Figure 26, they described charge injection by a network of infinitesimal small resistors representing the vertical resistance (denoted by r_y) and lateral bulk resistance (denoted by r_x). This model has been successfully used to investigate the contact effects not only of silicon-based transistors,^{237,238} but of OFETs as well, especially in the staggered structure.^{211,234,239,240} The model is based on a few assumptions. It is assumed that i) the mobility and resistivity are constant throughout the whole semiconductor film, ii) charge carriers are vertically injected from the contact and transported horizontally in the accumulated channel, and iii) the vertical current flows only in the contact area, i.e. in the region of the device defined by the gate/source and gate/drain overlap.

For the numerical analysis, r_x and r_y are defined as

$$r_x = R_{sh}dx/W \quad (9)$$

$$r_y = \rho_c/Wdx. \quad (10)$$

Here, R_{sh} is a sheet resistance (units of $\Omega/\text{sq.}$), and ρ_c is the specific contact resistivity ($\Omega \cdot \text{cm}^2$) equal to $R_c \times A_c$ with $A_c = WL_c$ (A_c is the contact area, W the contact (or channel) width, and L_c the contact length). $J_y(x)$ is the vertical current density at x . Using these definitions, one obtains from Figure 26

$$\frac{dI_x(x)}{dx} = J_y(x)W = \frac{V_x(x)}{\rho_c}W \quad (11)$$

$$\frac{dV_x(x)}{dx} = I_x(x) \frac{R_{sh}}{W} \quad (12)$$

With the boundary conditions of $I_x(0) = I_0$ and $I_x(-L_c) = 0$, Equations 11 and 12 can be solved for $V_x(x)$,

$$V_x(x) = I_0 L_T \frac{R_{sh}}{W} \frac{\cosh[(L_c - |x|)/L_T]}{\sinh(L_c/L_T)} \quad (13)$$

where $L_T = \sqrt{\rho_c/R_{sh}}$ is the characteristic length of injection denoted as the transfer length. It can be considered as the critical distance over which most of charges are injected from the metal contacts into the semiconductor.²⁴¹ At $x = 0$, the width-normalized source contact resistance r_S can be expressed as (R_S : source contact resistance)

$$r_S = R_S \cdot W = L_T R_{sh} \coth\left(\frac{L_c}{L_T}\right) \quad (14)$$

Under the assumption that ρ_c is identical for source and drain electrodes region, we obtain $R_c = 2R_S$.

Recently, Ante et al. reported an analysis of the influence of R_c and L_c on the effective

mobility and the operating frequency in staggered OFETs based on the current crowding model.²¹¹ They aggressively scaled the contact and channel dimension down to a few μm and investigated the relationship between L_T , L_c , and R_c .

In Figure 27, effective mobilities of DNTT OFETs and width-normalized R_c are plotted as a function of L_c in (a) and (b), respectively.²¹¹ Reducing L_c below L_T leads to a significant increase in R_c and thus to a decrease in the effective mobility. For best fitting results, an extended contact length L_{ext} was added to the contact length L_c in order to calculate the contact resistance from Equation 14. The extended contact length L_{ext} is motivated by the fact that charge is not only injected vertically at these small dimensions. At the edges of the overlap area between source electrode and gate, charge is injected in a nonperpendicular direction as well due to the electrostatic potential distribution in the organic semiconductor, which was found in a two-dimensional numerical model simulations.²³³

Figures 27(c) and (d) show results of a simulation examining the influence of the contact resistance on the maximum switching frequency of OFETs. The cutoff frequency of staggered OFETs is plotted for varying L and L_c assuming that $R_c = 0$ (c) or that R_c is a function of L_c described by Equation 14 (d). In the ideal case of zero R_c , the cutoff frequency increases dramatically into the gigahertz range by reducing L_c and L . However, when contact effects induced by the variation of L_c and L are taken into account, the cutoff frequency is found to be strongly limited. According to the dynamic performance of the actual OFET devices, indeed, the measured switching frequency shows reasonable agreement with the simulation results in Figure 27(d). This indicates that R_c of OFETs plays a significant role limiting the effective mobility as well as the operating frequency. Therefore, contact doping which can reduce R_c efficiently is a promising strategy to improve the operating frequency beyond the megahertz level.

To extract the contact resistance R_c experimentally, the transfer (or transmission) line method (TLM) is commonly used in the field of OFETs,^{242,243} even though it was first

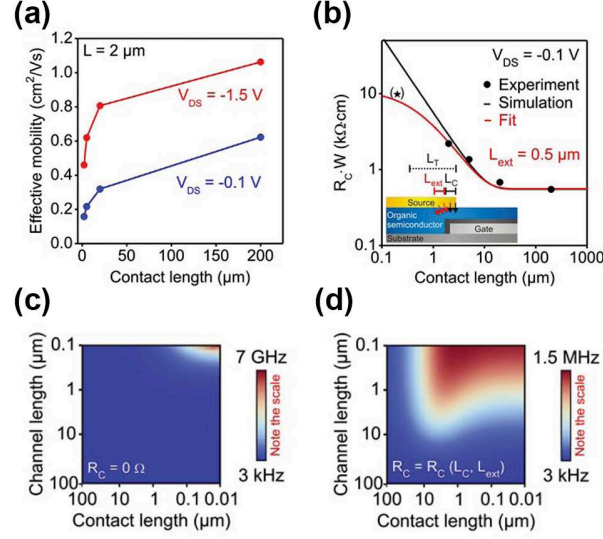


Figure 27: (a) Effective mobilities of DNTT OFETs as a function of L_c in the saturation (red curve) and the linear regimes (blue curve) ($L = 2\mu\text{m}$). (b) Width-normalized R_c as a function of L_c . Simulation results of cutoff frequency depending on L and L_c of staggered OFETs assuming that $R_c = 0$ (c) or that R_c depends on L_c (d). Adapted with permission from²¹¹. Copyright 2012 John Wiley and Sons.

developed for amorphous silicon thin film transistors (TFTs).²⁴⁴ The TLM method relies on the approximation that the contact resistance does not depend on the channel length L , but is mostly governed by the interfacial properties, whereas R_{ch} scales with the channel length L .

From the drain current I_D vs. voltage V_{DS} relation in the linear regime, the channel resistance R_{ch} can be determined by

$$R_{ch} = \left. \frac{\partial V_{DS}}{\partial I_D} \right|_{V_{DS} \rightarrow 0} = \frac{L}{\mu C_{ins} W (V_G - V_T)} \quad (15)$$

where μ denotes the field-effect mobility, C_{ins} the capacitance of the dielectric per unit area, V_G the gate voltage, and V_T the threshold voltage. Hence, if the total resistance R_{tot} is plotted vs. the channel length L at a given V_G , a linear relationship is obtained. According to Equations 8 and 15 the y -axis intercept can be interpreted as the contact resistance R_c . This relation can be used as well to obtain other parameters such as the charge carrier mobility μ and the threshold voltage V_T by varying V_G .²⁴⁵ In Figure 28, an

representative example for the TLM technique is shown.²⁴³ Here, the TLM can be used to determine the dependency of R_c on V_G from the y -axis intercept variation for varying V_G .

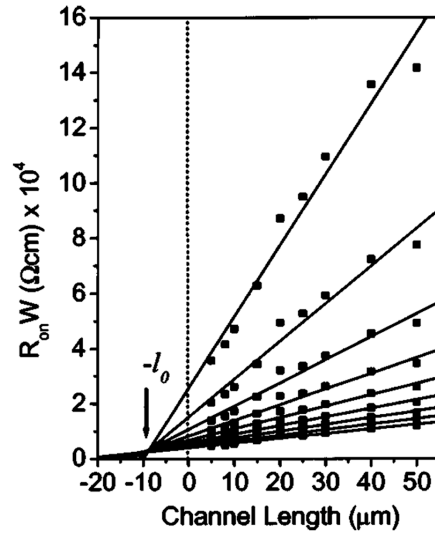


Figure 28: Normalized device resistance as a function of L for gate voltages V_G ranging from -20 to -100 V, extracted from pentacene transistors with L between 5 and 50 μm and a channel width of 400 μm . The y -intercepts of the solid lines give the contact resistance R_c at varying V_G . Reprinted from²⁴³, with the permission of AIP Publishing.

Although the TLM technique is often used to extract the contact resistance R_c and is straightforward to interpret, it shows some major drawbacks, which are in particular caused by the implicit assumption that the device parameters are identical for different samples with varying channel length L . However, sample-to-sample variations are hard to avoid in the experiment, especially if the transistors are not based on single-crystalline materials. In these thin-film transistors, the electrical performance is sensitive to morphological changes in the organic films as well as small changes in the fabrication conditions, making it difficult to reduce a significant spread in device parameters.²⁴⁶ To overcome these problems, Xu et al. proposed to modify the conventional TLM by plotting R_{tot}/L versus $1/L$ instead of R_{tot} vs. L . Thus, the slope of the fitted line becomes a function of R_c , which generally shows less variation compared to the channel-related parameters, which leads to more reliable fitting results.²⁴⁷

The gated four-point-probe (gFPP) technique is another measurement method that

is used to directly determine the contact resistance of a single OFET device. Originally, it was proposed for inorganic TFTs and was used to measure the intrinsic electrical performance,²⁴⁸ but it has been extended to quantify the contact resistance R_c as well.^{240,249,250} The fundamental circuit diagram and the principle of the gFPP technique are schematically illustrated in Figure 29.

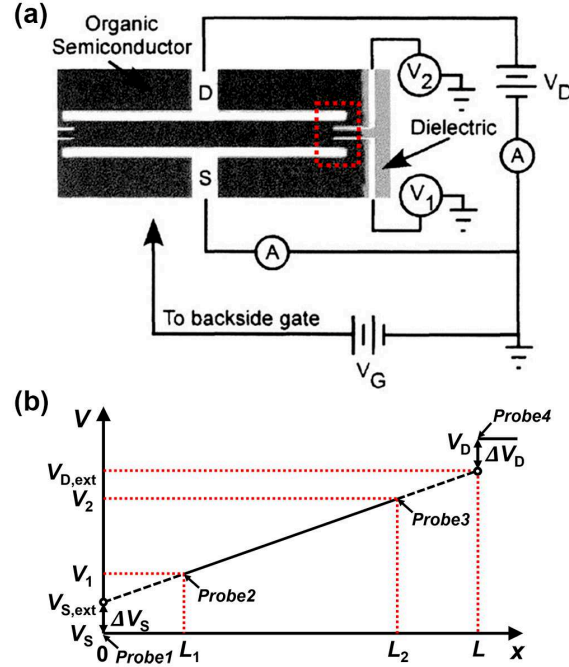


Figure 29: (a) Sketch of the gated four-point-probe (gFPP) measurement configuration. Reproduced from²⁴⁹, with the permission of AIP Publishing. (b) Principle of gFPP. The schematic represents the area enclosed by the red dotted line in (a).

Two additional high-impedance probes are defined in the channel between the source and drain to measure the local potential inside the channel. As shown in Figure 29, the channel potential at the source and drain electrode $V_{S,ext}$ and $V_{D,ext}$, i.e. at $x = 0$ and $x = L$, can be extrapolated from the two potentials V_1 and V_2 at the position of the two additional probes located at $x = L_1$ and $x = L_2$. From the difference between the externally applied potential, i.e. V_S and V_D , and the extrapolated potential, the voltage drops ΔV_S and ΔV_D at both contacts are extracted, which can be converted to the contact resistance at source R_S and drain R_D electrodes by Ohm's law ($R_S = \Delta V_S / I_{DS}$).

The gFPP method is only exact if a constant mobility is assumed, which is usually not fulfilled at high drain voltages or low gate voltages, i.e. when the device is in the saturation regime.²⁴⁹ Furthermore, the gFPP technique requires the deposition of two additional electrodes, which adds to the device complexity. Still, due to its ability to extract R_S and R_D directly from measurements on a single device it is highly beneficial for the study of contact effects at the metal/organic interface as recently reviewed by Natali and Caironi.²⁴⁵

5.1.2 Quasi-Ohmic Contacts Realized by Doping

Doping primarily creates free charges by a charge transfer between the host and the dopant molecules, which leads to an increase of electrical conductivity by several orders of magnitude in the bulk.³³ However, at the organic/metal interface, doping can have an even more pronounced effect on the electrical performance of organic electronic devices.

As shown in Figure 25, a space charge layer is formed at the metal/organic interface. Without doping, the width of the space charge region (W_{sc}) is relatively large (e.g. >10 nm), and charges are injected by thermionic emission across the energy barrier. Therefore, unless a metal contact is chosen, whose work function is similar to the LUMO level of organic semiconductor (for electron injection) or similar to the HOMO (for hole injection), these energy barriers greatly hinder the efficient charge injection.

Doping can help to overcome this limitation. In 2001, Blochwitz et al. investigated the level alignment at the interface between indium tin oxide (ITO) and either intrinsic ZnPc or ZnPc doped with F_4 -TCNQ by UPS and XPS.¹² Figure 30 shows the energy level scheme obtained from the experiment. It can be observed that the HOMO level of doped ZnPc is shifted towards the Fermi level and the level bending becomes stronger compared to the intrinsic layer. This observation agrees well with the Fermi Level shift observed in doped pentacene shown in Figure 6 (Section 4).

Although the hole injection barrier is almost identical for the doped and intrinsic case

(about 1.2 eV), the thickness of the the space charge region shrinks to less than 5 nm in the doped case. This very thin space charge layer allows for effective charge injection by tunneling, which was not achieved in the intrinsic layer. Hence, a quasi-ohmic contact known from inorganic semiconductors is formed, which ensures efficient injection into the semiconductor.

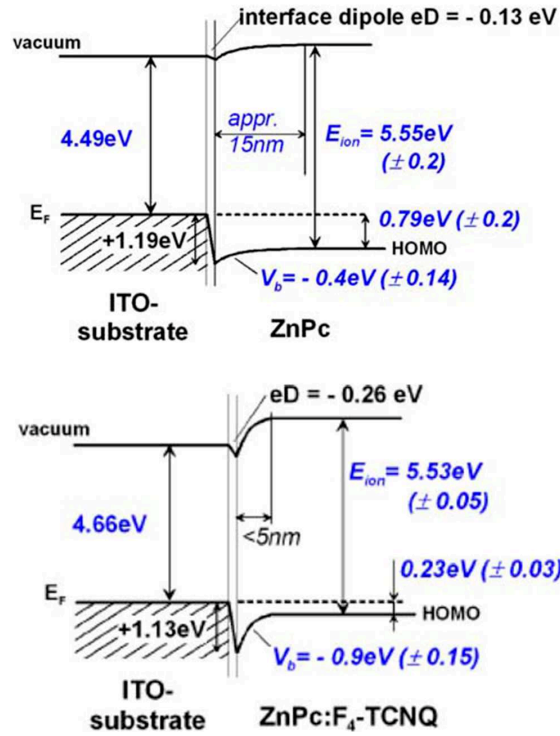


Figure 30: Energy level diagrams derived from UPS and XPS for undoped (top) and F₄-TCNQ-doped (bottom) ZnPc at the contact to ITO. Reprinted with permission from³³. Copyright 2003 Elsevier.

This mechanism was studied for a wide range of electrode/organic semiconductor combinations as shown in Figure 31.⁵⁴ For all cases shown in Figure 31, doping leads to strong level bending at the interface between contact and organic layer. The space charge layer is approximately 5 nm thick, which can be easily tunneled through. Hence, the interfacial barrier height ϕ_0 only plays a minor role for charge injection, i.e. charge injection is not limited by a mismatch between the metal work function and the organic transport level.

The influence of quasi-ohmic contacts realized by doping is shown in Figure 32. As

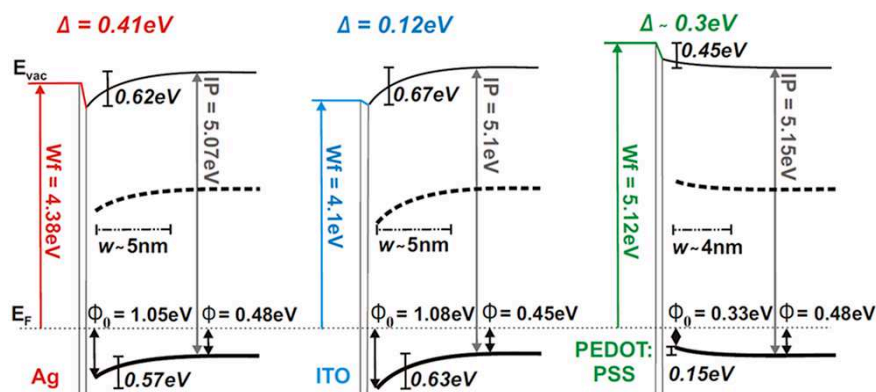


Figure 31: Energy level diagrams obtained from UPS and XPS measurements for F_4 -TCNQ-doped MeO-TPD at the interfaces on the different electrodes (Ag, ITO, and PEDOT:PSS). Reprinted from ⁵⁴, with the permission of AIP Publishing.

seen from the J-V characteristics of a metal-organic semiconductor-metal structure, using doped layers (F_4 -TCNQ-doped N,N' -diphenyl- N,N' -bis(1-naphthyl)-1,1'-biphenyl-4,4'-diamine (α -NPD) or alpha-sexithiophene (α -6T)) leads to a linear, i.e. ohmic relationship between current and voltage.²⁵¹ Therefore, contact doping generating such quasi-Ohmic contacts can be effectively applied to minimize contact resistances and to develop high performance OFETs.

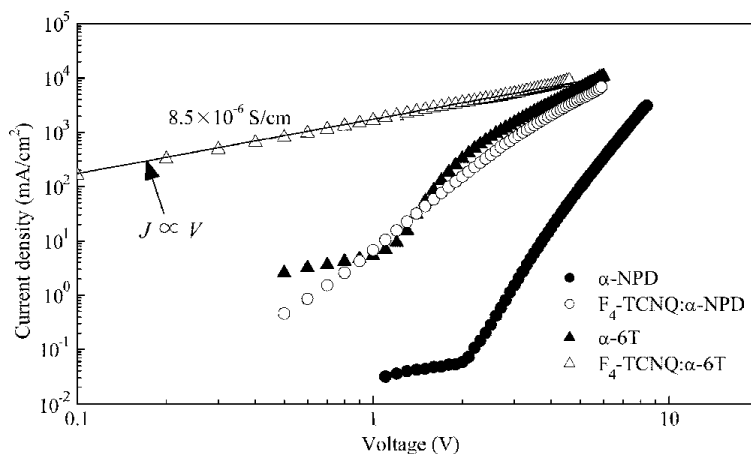


Figure 32: Current density-voltage characteristics of hole only transport devices, i.e. ITO/organic layer/Ag/Al, with α -NPD, F_4 -TCNQ-doped α -NPD, α -6T, and F_4 -TCNQ-doped α -6T layers. Reprinted from ²⁵¹, with the permission of AIP Publishing.

5.1.3 Contact Doping for P-Type Transistors

As discussed so far, the main goal of contact doping is to reduce the contact resistance and accordingly to enhance charge injection in OFETs. The most straightforward way to realize a doped organic contact is to insert a pure dopant layer, e.g. a layer of metal oxides,^{227,252–255} metal halides,^{15,20,256,257} or molecular dopants,^{13,207,258} between the metal and the organic semiconductor.

Ante et al. studied the effect of a pure dopant layer on the contact resistance of p-type OFETs with submicrometer channel length (see Figure 33).¹³ The OFETs consist of a 5.7 nm-thick AlO_x gate dielectric, an organic SAM grown on top of the oxide layer, a layer of the organic semiconductor DNTT, and gold source/drain contacts. For a contact doping, a pure layer of the proprietary dopant NDP-9¹ is inserted between the metal and the organic layer. The source/drain contacts are structured by a suspended resist bridge defined by electron-beam lithography used as shadow mask, which yields nanoscale channel lengths.¹⁴

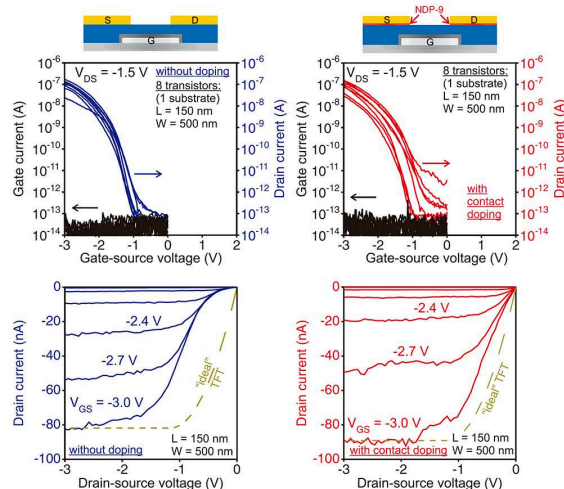


Figure 33: Device structure and transfer/output characteristics of submicrometer OFETs without (left) and with (right) contact doping. Reprinted with permission from¹³. Copyright 2011 John Wiley and Sons.

As clearly seen in the output curves of Figure 33, contact doping with NDP-9 not

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only increases the drain current but as well diminishes the nonlinearity at a low voltages compared to the undoped case. For example, the drain current at $V_{GS} = -3$ V and $V_{DS} = -0.5$ V increases from -5 nA to -30 nA. This increase of the drain current is obviously attributed to a reduction in the contact resistance at the contact/semiconductor interfaces by contact doping.

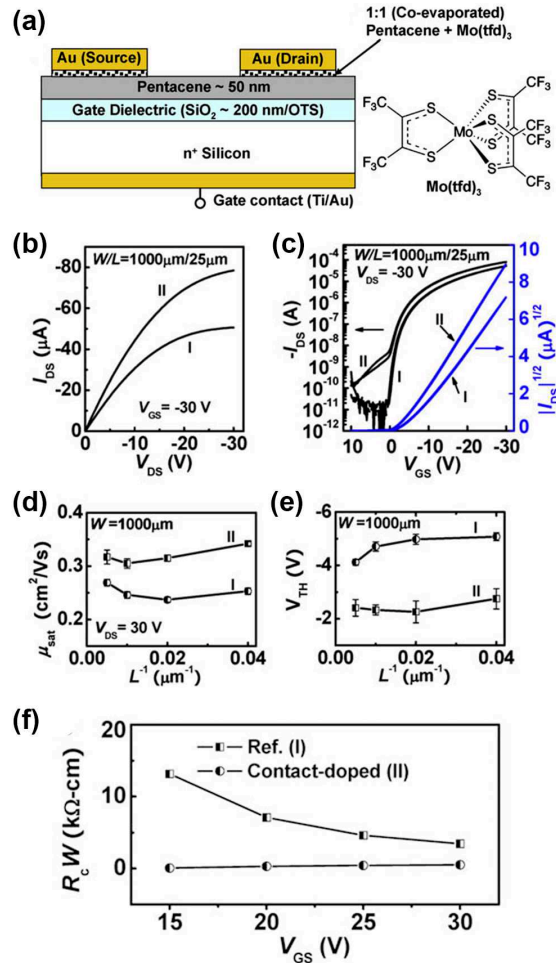


Figure 34: (a) Device structure and chemical structure of Mo(tfd)₃. (b) The output and (c) the transfer characteristics of OFETs without and with contact doping. Comparison of (d) the mobility, (e) the threshold voltage, and (f) the contact resistance. Adapted with permission from ¹⁶. Copyright 2010 Elsevier.

Another viable strategy to realize efficient contact doping is the co-deposition of host and dopant materials. This process has been extensively investigated either in solution processed films ^{162,166,213} or in films prepared by co-evaporation. ^{21,188,259–261} A

representative study is summarized in Figure 34.¹⁶ Tiwari et al. used Mo(tfd)₃ as p-dopant for pentacene OFETs. As mentioned in Section 3, the electron affinity of Mo(tfd)₃ was reported as 5.6 eV,⁹⁶ which is large enough to dope pentacene. For contact doping, a 10 nm thick layer of a 1:1 mixture of Mo(tfd)₃ and pentacene is selectively deposited in the source/drain electrodes area (see Figure 34(a) for a sketch of the structure). From the electrical characteristics, it is seen that I_D increases about 50% by contact doping for devices with $W/L = 1000 \mu\text{m}/25 \mu\text{m}$. Figures 34(d) and (e) show the improvement of the effective mobility and the threshold voltage compared to the undoped OFETs regardless of the channel length. To examine the effect of doping more specifically, the contact resistance for each type of devices are extracted using TLM as discussed above. The plots of width-normalized contact resistance $R_c W$ at different V_{GS} are compared in Figure 34(f), and it is found that the contact resistance is indeed decreased, for example, from 3.4 k Ω -cm to 0.5 k Ω -cm in the strong accumulation regime ($V_{GS} = -30 \text{ V}$).

OFETs using the staggered structure usually show a lower contact resistance than coplanar OFETs, which is due to the relatively large charge injection area. From an industrial point of view, however, bottom gate and bottom contact configurations offer the possibility to employ high resolution lithography to structure devices without damaging the organic semiconductor. To take advantage of this configuration, the doped injection layer has to be structured by lithography as well to avoid unwanted channel doping and a decrease in the ON/OFF ratio of the transistor.

Nicht et al. demonstrated a promising approach that combines the properties of SAMs, which selectively attach to gold electrodes, with molecular doping to improve charge injection into coplanar OFETs (cf. Figure 35).²⁶² In this study, F₄-TCNQ is functionalized with sulfur-containing triazole anchor groups via nucleophilic substitution in a simple substitution reaction. By selective assembly of a sulfur-functionalized F₄-TCNQ derivative exclusively on the source and drain electrodes, current injection is significantly increased in pentacene-based OFETs. The effective mobility is also improved by a factor of three

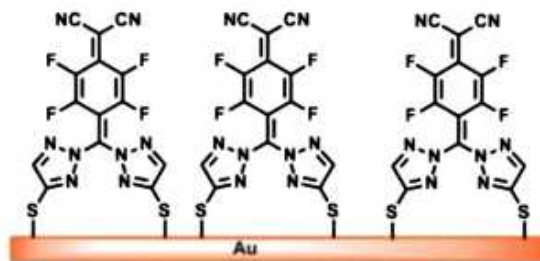


Figure 35: Functionalizing organic dopants by groups that selectively bind to metals allows for the directed assembly of self-assembled monolayers of dopants on source and drain electrodes. Nicht et al. proposed a sulfur-functionalized F₄-TCNQ derivative to increase injection into bottom contact OFETs. Reprinted with permission from²⁶². Copyright 2014 Elsevier.

from 0.08 cm²/Vs to 0.24 cm²/Vs, which indicates that a SAM of functionalized dopant is effective for the reduction of contact resistance.

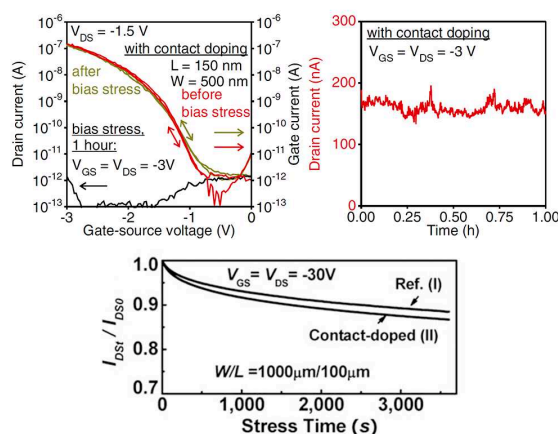


Figure 36: Bias-stress measurement for contact-doped OFETs. Transfer characteristics of a submicrometer OFET with contact doping before and after bias stress, and the measured drain current during bias stressing time (top). Reprinted with permission from¹³. Copyright 2011 John Wiley and Sons. Decay of drain current measured continuously over a period of 1 h (bottom). Reprinted with permission from¹⁶. Copyright 2010 Elsevier.

The morphological stability of doped films is a major concern of doping in particular for elemental dopants (see as well Section 4). If dopants are mobile in the film there is a possibility of dopants to diffuse into the channel, which would lead to unintentional doping of the channel and an increase in the off-state current. However, Figure 36 shows that molecular dopants and transition metals show sufficient stability in the

transistors. The results of bias stress measurements for NDP-9-doped submicrometer OFETs¹³ and Mo(tfd)₃-doped pentacene OFETs¹⁶ are summarized. In both cases, the electrical properties are not significantly changed for increasing stressing times, which indicate that contact doping is effective without compromising the electrical stability.

5.1.4 Contact Doping for N-Type Transistors

As already mentioned in Section 4, n-doping is rather challenging compared to p-doping due to the increased sensitivity of n-dopants to air exposure. However, in order to design complementary circuits, the development of n-channel OFETs with a performance comparable to that of p-channel OFETs is essential. Therefore, similar to p-contact doping, a wide range of approaches have been proposed to induce n-doping, leading to the improvement of charge injection of n-type transistors. For example, interlayers of Ca,²⁶³ TiO_x,²²⁹ conjugated polyelectrolyte,²⁶⁴ or water-soluble polyfluorene derivatives²⁶⁵ were introduced between organic semiconductor and electrodes and soluble mixtures such as CsF-doped PCBM,²¹³ P(NDI2OD-T2) mixed with CoCp₂ or CsF,²⁶⁶ and C₆₀, C₇₀, or PCBM doped by N-DMBI were shown to improve injection as well.²⁶⁷

Recently, Singh et al. demonstrated the reduction of the contact resistance by selective contact doping in bottom-gate and top-contact C₆₀ OFETs using rhodocene dimer as a dopant (Figure 37(a)).¹⁷ In Figures 37(b) and (c), the transfer and the output curves of contact-doped and undoped OFETs with a channel geometry of $W/L = 1200\ \mu\text{m}/25\ \mu\text{m}$ are plotted, respectively. The output characteristics at $V_{GS}=24\ \text{V}$ shows that contact doping leads to a three fold increase in I_D compared to the reference devices. Note that the off-current remains unchanged after doping, which is an indication that only the contact area is doped. As shown in Figure 37(d), the mobility decreases for both doped and reference devices as the channel length reduces, but this decrease is more pronounced for the undoped case. The average mobility values for the undoped and the doped devices are extracted as $0.48\ \text{cm}^2/\text{Vs}$ and $1.65\ \text{cm}^2/\text{Vs}$ for $L = 25\ \mu\text{m}$, i.e. the

mobility is enhanced by about a factor of 3 by contact doping.

Singh et al. studied as well the effect of doping on the contact resistance using the transmission line method (TLM). The results indeed confirm that the contact resistance is reduced from 9.7 k Ω -cm of reference OFETs to 5.5 k Ω -cm of contact-doped devices at $V_{GS} = 24$ V (Figure 37(e))

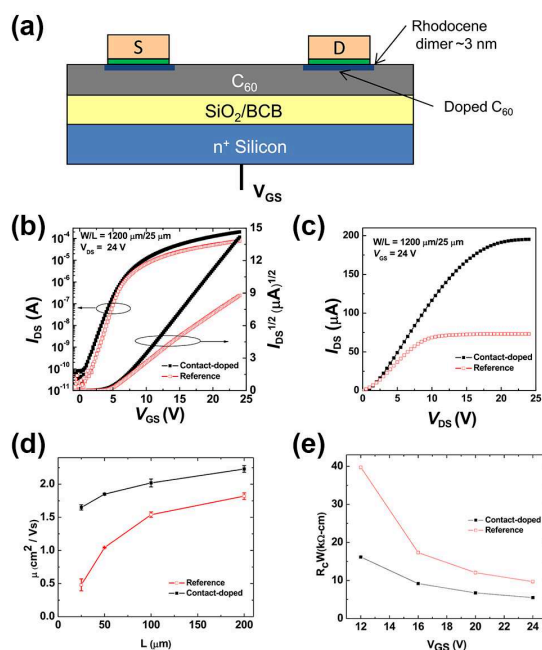


Figure 37: (a) Device structure of contact-doped C₆₀ OFETs. (b) Transfer and (c) output characteristics of OFETs without and with contact doping. Comparison of (d) the mobility and (e) the contact resistance at varying channel lengths. Reproduced from¹⁷, with the permission of AIP Publishing.

The influence of n-doping on the air-stability of n-channel OFETs was studied by Oh et al.²⁶⁸ In this study, an active layer, N,N'-dibutyl-1,7-difluoroperylene-3,4:9,10-tetracarboxylic diimide (F-PTCDI-C4) is doped by pyronin B (see Figure 38(a)). Doping is introduced either in the contact region, in the whole bulk of the organic semiconductor, or selectively at the source/drain contacts.

In Figure 38(b), the transfer characteristics of doped and undoped F-PTCDI-C4 OFETs measured in N₂ atmosphere are plotted. Field-effect mobilities of 0.2 to 0.3 cm²/Vs are obtained with on/off ratios of 10⁶, regardless of the position of the doped layer in

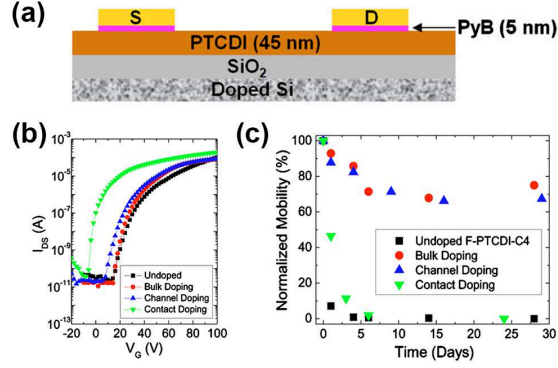


Figure 38: (a) Device structure of contact-doped F-PTCDI-C₄ OFET. (b) Transfer characteristics of OFETs without and with n-doping. (c) Comparison of the normalized mobility. Reproduced from²⁶⁸, with the permission of AIP Publishing.

the device. However, compared to undoped, bulk doped, or channel doped devices, the threshold voltage of contact doped OFETs shift towards more negative voltages, indicating that the devices are switched on at lower voltages. The air-stability of pyronin B-doped F-PTCDI-C₄ OFETs is also investigated by periodically measuring the electrical performance while the devices are being stored in air. In Figure 38(c), the extracted mobilities are normalized by those initially obtained in N₂ atmosphere for comparison. Interestingly, the performance of undoped F-PTCDI-C₄ OFETs exhibit an immediate degradation by air-exposure, and the contact-doped devices gradually degrade as well. This means that contact doping leads to the enhancement of charge injection, but charge trapping in the channel region by air-exposure still remains a major issue in this particular study. In contrast, the air-stability of channel-doped F-PTCDI-C₄ OFETs is improved significantly. The detailed underlying principles are discussed in Section 5.2.

Contact doping was as well shown to be effective for solution-processed OFETs. A recent study argued that PCBM based transistors can be n-doped by cesium fluoride (CsF) by either solution blending or inkjet printing. Ohmic injection independent of the gate bias is observed.²¹³ Particularly, in this study, the selective tuning of the operational mode from ambipolar to unipolar transport is achieved by molecular doping, which will be discussed in detail Section 5.2.3. Here, PCBM and a perfluoropolymer, namely CYTOP,

are utilized as the pristine organic semiconductor and the gate dielectric, respectively, and top-gate/bottom-contact (TG/BC) structure is adopted to investigate the doping effect on the charge injection and transport. In this system, as the Fermi level of CsF (1.9 ~ 2.0 eV) lies above the LUMO level of PCBM (≈ 3.7 eV), electron transfer from CsF into PCBM is energetically favorable, leading to efficient n-doping.

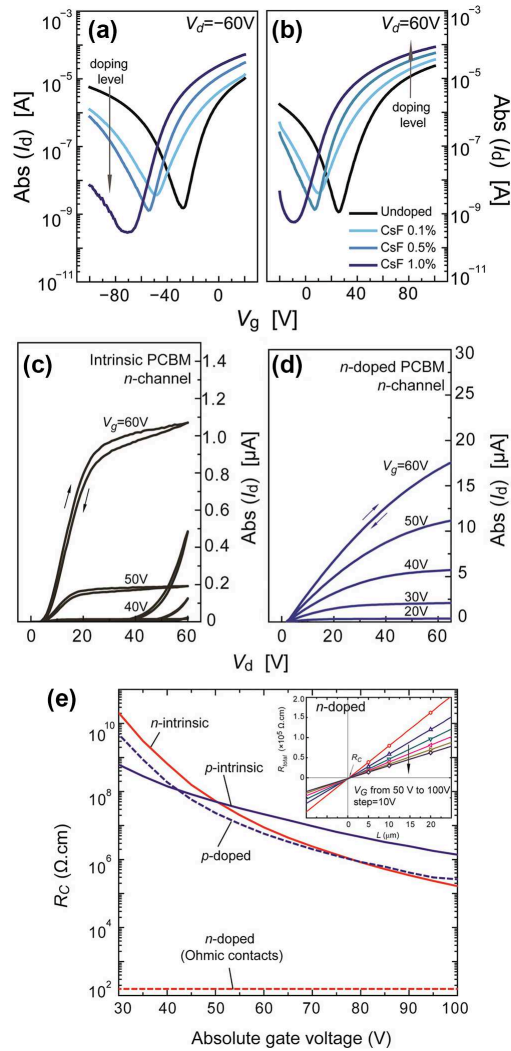


Figure 39: Transfer characteristics of CsF-doped PCBM OFETs for (a) p-channel and (b) n-channel operation. Output characteristics of (c) undoped and (d) doped PCBM OFETs for n-channel operation. (e) R_c of TG/BC PCBM OFETs with varied gate-voltage. Inset shows R_c of CsF-doped PCBM OFETs extracted by modified-TLM. Adapted with permission from ²¹³. Copyright 2014 John Wiley and Sons.

Figures 39(a) and (b) show the transfer characteristics of CsF-doped PCBM OFETs for

p-channel and n-channel operation. It is clearly seen that the undoped devices originally show ambipolar charge transport, however p-channel operation is increasingly suppressed while n-channel operation becomes dominant as the n-doping concentration increases. This conversion from ambipolar to unipolar feature is mainly attributed to the Fermi level shift towards the transport level by doping. For 1.0 wt.% CsF-doped OFETs, the electron mobility and the hole mobility are remarkably changed from $0.067 \text{ cm}^2/\text{Vs}$ and $0.030 \text{ cm}^2/\text{Vs}$ to $0.13 \text{ cm}^2/\text{Vs}$ and $5.9 \times 10^{-5} \text{ cm}^2/\text{Vs}$, respectively, indicating negligible hole transport.

The output curves of undoped and n-doped PCBM OFETs are plotted in Figures 39(c) and (d). The nonlinearity at low V_d originates from large charge injection barriers for undoped devices, but disappears after n-doping, which confirms that the contact resistance is reduced. The contact resistance is extracted using modified-TLM introduced above (Figure 39(e)). It turns out that for n-type doping the contact resistance is dramatically reduced, and becomes independent of the gate voltage, i.e. exhibiting an Ohmic contact behavior. This is a strong evidence that CsF doping reduces the thickness of the depletion layer and enables tunneling injection at the interface of electrode and doped PCBM.

5.2 Channel Doping

Although contact doping is much more common, doping the channel region of the transistor receives increasing attention as well. Channel doping has a couple of advantages; most importantly it presents a way to fine tune the transistor characteristic, to control the threshold voltage, and to increase the stability of the transistors.

5.2.1 Doped Organic MIS Junctions

A doped metal-insulator-semiconductor (MIS) junction is at the heart of every doped organic transistor. The band structure of a p-doped MIS junction is shown in Figure 40. Here, W_f denotes the work-function of the metal, IP the ionization potential of the

organic semiconductor, η the difference between the ionization potential and the Fermi level in the organic semiconductor, and δ a dipole located at the interface between the insulator and the semiconductor.

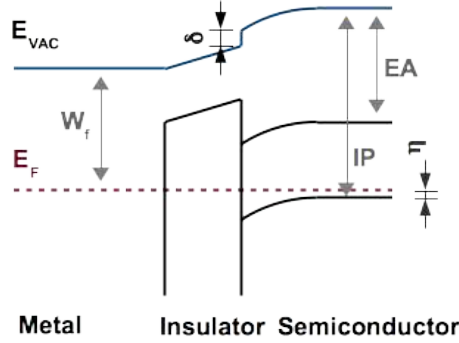


Figure 40: Alignment of energy levels in a metal-insulator-p-doped semiconductor junction. Due to the difference in chemical potential of the metal and semiconductor, a thin layer at the insulator-semiconductor interface is depleted forming a negative space charge layer. The thickness of this depletion layer can be modulated by the voltage applied between the metal and the semiconductor.

Similar to metal/semiconductor junctions discussed previously (Section 5.1) charge is redistributed inside the junction until the work function of the metal and the organic semiconductor align and a constant Fermi Level is formed. For example, if the work function of the semiconductor exceeds the work function of the metal (i.e. $IP - \eta > W_f$), the p-doped layer is depleted at the interface to the insulator, which generates a negative space charge layer.⁵⁴ The thickness of this depletion layer d_{dep} can be calculated as^{68,269}

$$d_{dep} = \frac{\epsilon_0 \epsilon_s}{C_{ins}} \left[\sqrt{1 + \frac{2C_{ins}^2 (V - V_{FB})}{eN_A \epsilon_0 \epsilon_s}} - 1 \right] \quad (16)$$

with ϵ_0 the permittivity of vacuum, ϵ_s the relative permittivity of the semiconductor, C_{ins} the capacitance of the insulator per unit area, $e = +1.6 \times 10^{-19}$ C the elementary charge, V the voltage applied between metal and semiconductor (counting the potential on the metal positive), and N_A the doping concentration in the p-doped semiconductor.

V_{FB} denotes the so called flatband voltage, which is defined as the voltage that has to

be applied between the organic semiconductor and the metal to induce flat energy levels. Including charged trap states Q_{trap} at the insulator/organic semiconductor interface, in the insulator, or in the organic semiconductor, the flatband voltage becomes

$$eV_{FB} = W_f - (IP - \eta) - \delta - Q_{trap}/C_{ins} \quad (17)$$

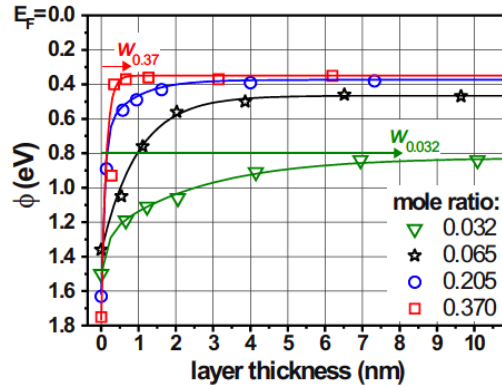


Figure 41: Evolution of the HOMO level (Φ) of an organic MIS junction consisting of a silver cathode, an intrinsic layer of MeO-TPD and a p-doped layer of MeO-TPD of varied thickness. From the plot, the thickness of the depletion layer and the flatband voltage can be deduced, which shows the same trends as predicted by Equations 16 and 17. Reprinted from⁵⁴, with the permission of AIP Publishing.

The thickness of the depletion layer (Equation 16) and the flatband voltage (Equation 17) can be determined experimentally by ultraviolet photoelectron spectroscopy. Olthof et al. studied the alignment of energy levels in a MIS junction consisting of a silver electrode, 5 nm of the intrinsic organic semiconductor MeO-TPD operating as insulator, and a p-doped MeO-TPD layer of varying thickness.⁵⁴ The position of the HOMO level for different thickness of the doped layer is shown in Figure 41. As expected, the HOMO level bends upwards at the interface between the intrinsic and the doped MeO-TPD layer, indicating the formation of a depletion layer. Level bending becomes stronger for increased doping concentrations and, in accordance to Equation 16, the depletion layer shrinks with increasing doping concentrations. Olthof et al. found a depletion layer thickness of 8 nm for a molar doping ratio of 0.032, which decreases to 0.6 nm at a molar

doping ratio of 0.37. In the same way, the flatband voltage increases from 0.66 V (MR = 0.032) to 1.4 V (MR = 0.37), which can be explained by the shift of the Fermi-Level of the doped organic semiconductor toward the HOMO level, e.g. by a decrease of η in Equation 17.

5.2.2 Doped Organic Transistors

Tuning the MIS junction of an organic transistor by doping has profound influences on the transistor operation. The general setup of a p-doped transistor was already shown in Figure 24. The semiconductor layer is doped by a p-dopant, e.g. by F₄-TCNQ or F₆-TCNNQ. In some cases, doping is restricted to a thin layer (thickness d_{dop}) at the interface to the insulator. Finally, the source and drain contacts can be p-doped²¹ or a metal can be used that shows sufficient hole injection to ensure efficient hole injection.

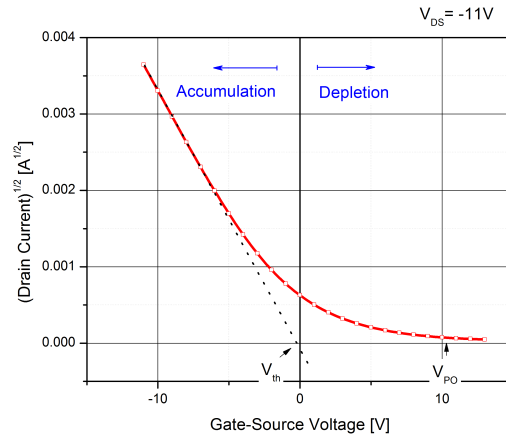


Figure 42: Transfer characteristic of a doped organic transistor as sketched in 24. The transistor consists of p-doped pentacene channel ($d_{dop} = 8$ nm, doping concentration 2 wt.%), intrinsic pentacene, and p-doped injection contacts. Details of the device structure and preparation conditions are given in.²¹

A representative transfer characteristic of a p-doped transistor is shown in Figure 42. As indicated in the Figure, the transistor can be operated in two regimes. At large negative voltages ($V_{GS} < V_{th}$) holes are electrostatically accumulated at the insulator/organic layer (accumulation regime), whereas at large positive voltages ($V_{GS} > V_{th}$) the doped channel

is increasingly depleted (depletion regime). To fully deplete the doped channel and switch the transistor off, the so called pinch-off voltage V_{PO} has to be applied to the gate. In the following, the accumulation and depletion regime are discussed separately.

Accumulation Regime If large negative voltages are applied to the gate of the p-doped transistor (i.e. for $V_{GS} < V_{th}$), holes are accumulated in the channel which increases the density of free holes beyond the doping concentration. Although the density of accumulated holes can significantly exceed the density of holes introduced by doping, doping has still a pronounced influence on the threshold voltage. This influence can be understood by considering the total density of free holes in the channel given by

$$p(x) = \frac{C_{ins}w (V(x) - (V_{GS} - V_{FB}))}{eA} + N_A \quad (18)$$

$$= \frac{C_{ins}w \left(V(x) - V_{GS} + V_{FB} + \frac{edN_A}{C_{ins}} \right)}{eA} \quad (19)$$

where w denotes the width of the channel, e the elementary charge, $A = dw$ the cross-section of the channel, d the thickness of the accumulation layer, and N_A the density of ionized dopants in the channel.

Using the relation $j = ep(x)E_x\mu_p$ (E_x : electric field in x-direction, μ_p : hole mobility) and integrating across the whole channel leads to the following relation for the drain current I_D ²⁶⁹

$$-I_D = \frac{\mu_p C_{ins}w}{L} \left[\left(V_{GS} - V_{FB} - \frac{edN_A}{C_{ins}} \right) V_{DS} - \frac{1}{2} V_{DS}^2 \right] \quad (20)$$

$$= \frac{\mu_p C_{ins}w}{L} \left[(V_{GS} - V_{th}) V_{DS} - \frac{1}{2} V_{DS}^2 \right] \quad (21)$$

Equation 21 resembles the relation for the drain current of standard, i.e. undoped transistors, except that the threshold voltage is altered. Under the assumption that the thickness of the doped layer d_{dop} is small, one obtains

$$V_{th} \approx V_{FB} + \frac{eN_A d_{dop}}{C_{ox}} \quad (22)$$

Hence, the threshold voltage depends on the doping concentration and the thickness of the doped channel and can be shifted by the doping concentration and thickness of the doped layer.

Equation 22 has been confirmed by several authors in the past. A shift of the threshold voltage due to doping was observed for a wide variety of material combinations including p- and n-type transistors and polymer and small molecular semiconductors. For example, transistors based on Pentacene doped by MoO_x ,²³ $\text{F}_4\text{-TCNQ}$ ²⁵⁹ or $\text{F}_6\text{-TCNNQ}$,^{21,212} TIPS-Pentacene doped by tris-[1-(trifluoroethanoyl)-2-(trifluoromethyl)ethane-1,2-dithiolene] ($\text{Mo}(\text{tfd-COCF}_3)_3$),²⁷⁰ P3HT doped by $\text{F}_4\text{-TCNQ}$,¹⁶⁶ perylene diimides doped by pyronin B,²⁶⁸ PCBM doped by N-DMBI,¹³² poly[4-(4,4-dihexadecyl-4H-cyclopenta[1,2-b:5,4-b']dithiophen-2-yl)-alt-[1,2,5]thiadiazolo-[3,4-c]pyridine] (PCDTPT) doped by I_2 vapor,²⁷¹ P3HT doped by gold nanoparticles,²⁷² or a mixed layer of sexithiophene and dihexyl-sexithiophene doped by oxygen²⁷³ show a shift of the threshold voltage with increasing doping concentration.

One representative example reported by the Heeger group is shown in Figure 43.²⁷¹ Lee et al. prepared polymer based transistors using the high-mobility material PCDTPT on top of a nanogrooved gate oxide to improve the alignment and order in the polymer film. The transistors are reported to show a charge carrier mobility of $18.1 \text{ cm}^2/\text{Vs}$ in the linear region and as high as $58.6 \text{ cm}^2/\text{Vs}$ in the saturation regime (Figure 43). Lee et al. argue that the as-prepared transistors show some unintentional doping leading to a

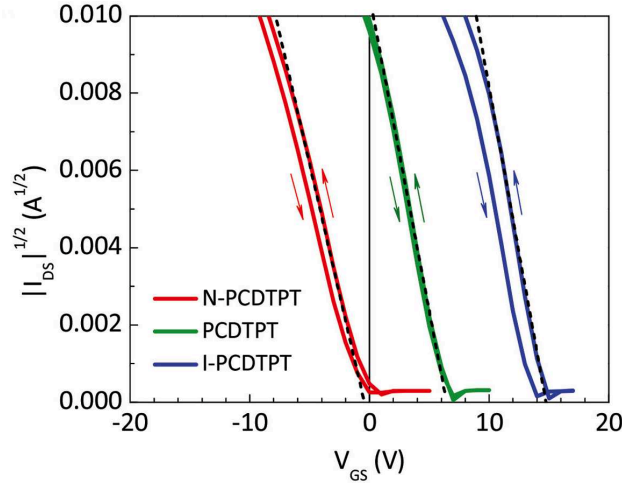


Figure 43: Doping the high mobility polymer PCDTPT leads to a strong shift in the transfer characteristic of the corresponding transistor.²⁷¹ The intrinsic transistor (green) shows a threshold voltage of around 7 V. P-doping by exposing the device to iodine vapor results in a shift toward higher voltages (blue), whereas exposing the device to ammonia vapor deactivates doping (red). Reprinted with permission from²⁷¹. Copyright 2016 John Wiley and Sons.

threshold voltage of approximately $V_{th} = 7$ V. Exposing the devices to iodine p-dopes the film further, resulting in an increase in threshold voltage to 14 V, whereas exposing the devices to ammonia vapor compensates the p-doping effect and a shift of the threshold voltage to close to 0 V.

Most reports on doped organic transistors study the influence of the doping concentration on the threshold voltages. However, according to Equation 22, the thickness of the doped layer can be used as well to tune the transfer characteristic. This effect was studied by the Leo group²¹ and the results are shown in Figure 44. Indeed, an increase in doped layer thickness d_{dop} increases the threshold voltage linearly.

Doping in organic transistors often comes at the expense of a reduced charge carrier mobility due to an increase in disorder in the film (see e.g. the discussion in Section 4).¹⁸⁵ To overcome this limitation, some researchers propose not to dope the organic semiconductor directly, i.e. not to mix a low concentration of dopants into the matrix semiconductor, but to add a dopant layer on top^{22,24,274} or below an intrinsic charge

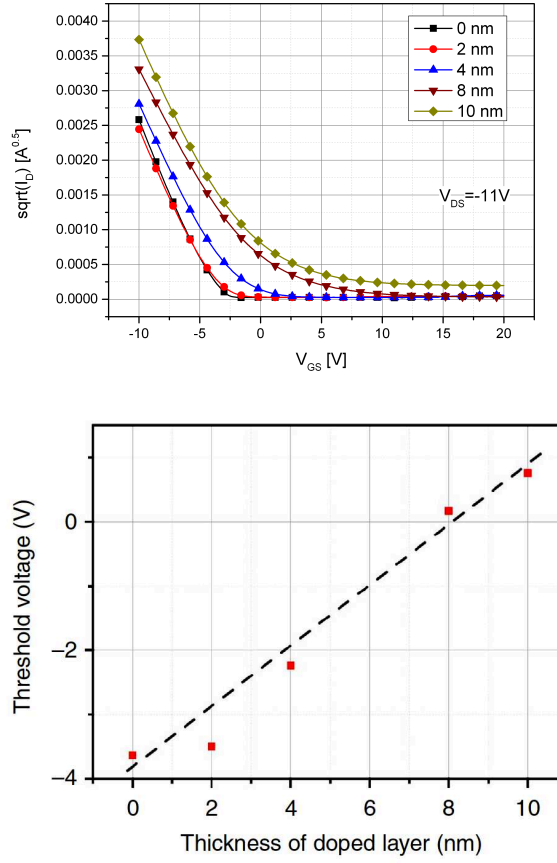


Figure 44: Influence of the thickness of the doped layer d_{dop} on the transfer characteristic of pentacene based transistors (top). As predicted by Equation 22 the threshold voltage increases for increasing d_{dop} (bottom). Reprinted from²¹. Copyright 2013 Nature Publishing Group.

transport layer instead,^{80,209,275–278} which is called surface, transfer, or remote²⁶⁰ doping.

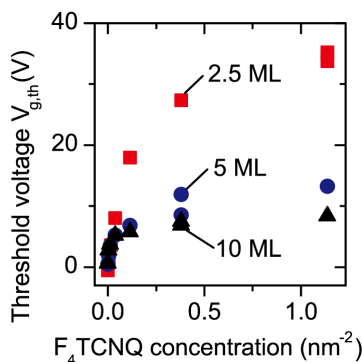


Figure 45: Transfer doping as proposed by Hählen et al.²⁴ Instead of doping the channel as in Figure 24, a pure F₄-TCNQ layer is deposited on top of an intrinsic pentacene layer. The strength of the shift of the threshold voltage depends on the thickness of the intrinsic pentacene layer. Reprinted from²⁴, with the permission of AIP Publishing.

For example, Hählen et al.²⁴ deposited a pure layer of F₄-TCNQ on top of a pentacene bottom-gate transistor and showed that the threshold voltage is increasing with increasing F₄-TCNQ coverage (Figure 45). However, this effect strongly depends on the thickness of the intrinsic pentacene layer, i.e. the shift of the threshold voltage for thicker pentacene layer is significantly reduced. Still, even at a thickness of 10 monolayers of pentacene a small shift in V_{th} is observed, which is rationalized by the grain structure of pentacene films and the ability of F₄-TCNQ to accumulate at grain boundaries.

The group of Kahn et al. proposed a similar approach to remotely dope pentacene films.²⁶⁰ They add a doped layer of α -NPD on top of an intrinsic layer of pentacene. The α -NPD film is reported to transfer some of the holes generated by doping to the intrinsic pentacene. Hence, the conductivity of the pentacene film increases and the threshold voltage shifts toward positive voltages.²⁶⁰

Self-assembled monolayers are another example of successful surface doping. Adding a self-assembled monolayer on top of the gate insulator is known to have a strong influence on the threshold voltage.^{279–281} Often, this effect is explained by the molecular dipole introduced at the insulator/organic interface δ , which is supposed to shift the flatband voltage according to Equation 17. Indeed, it was found that the shift of the

threshold voltage scales with the dipole moment of the SAM.^{282,283} However, it has been argued that interface dipoles can only explain small shifts in threshold voltage (up to a few volts)²⁸⁴ and that space charge layers induced by electronic coupling between the SAM and the organic semiconductor have to be taken into account in order to explain the observed threshold voltage shift.^{277,285}

A very successful example of an interfacial SAM layer that introduces additional space charge at the interface and operates as a local doping layer was studied by the group of Zojer et al.^{80,275,276} They functionalized the SAM with acidic end groups, which protonate the organic semiconductor at the interface (see as well Section 3.1.3 for a detailed discussion of the doping mechanism). The SAM hence creates a local doped channel, which shifts the threshold voltage according to Equation 22.

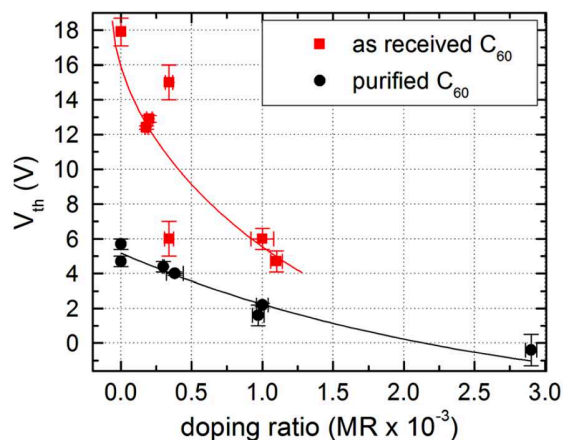


Figure 46: Filling traps in unpurified C₆₀ by the n-dopant [RuCp*(mes)]₂ leads to a strong shift of the threshold voltage. Reprinted from⁵¹, with the permission of AIP Publishing.

Doping can not only shift the threshold voltage by a generation of free charge carriers, but can as well be used to fill traps at the interface between gate insulator and organic semiconductor or in the organic semiconductor itself.⁸⁹ Figure 46 shows a recent example of this effect reported by Olthof et al.⁵¹ Small amounts of the n-dopant [RuCp*(mes)]₂ (see Section 3.2.7) were added to either highly purified or as received C₆₀. The corresponding shift of the threshold voltage is shown in Figure 46. The threshold voltage shift is much

stronger for the unpurified material compared to the purified C₆₀, which was interpreted as indication that doping mainly fills trap states in the device.⁵¹

Filling of matrix trap states by doping has the advantage that free charge carriers are not necessarily generated and that the ON/OFF ratio is not compromised by doping. This effect was used by Soeda et al.,²⁰⁹ who used F₄-TCNQ to p-dope the organic semiconductor C8-BTBT. Although the EA of F₄-TCNQ is too small to generate free holes on the C8-BTBT matrix, the threshold voltage of doped transistors is reduced to almost zero volts and the hysteresis of doped transistors is significantly reduced.²⁰⁹

Passivation of traps in doped organic layers was furthermore used by Wei et al.²⁵ to improve the stability of n-type transistors. Doping the organic semiconductor PCBM with DMBI-derivatives¹²⁹ (see Section 3.2.5) leads to a strong increase in air stability, which was rationalized in terms of a compensation of O₂ induced traps in PCBM. P-doping has a positive influence on stability as well. As observed by Hein et al., p-doping the channel region can lead to a reduction in gate bias stress effects in pentacene based transistors.²⁶ Similarly, Noh et al. found an improved air stability and smaller gate bias stress in doped ambipolar transistors.²¹³

Depletion Regime Doped organic transistors can as well be operated in the depletion region, e.g. at $V_{GS} > V_{th}$ (see Figure 42). No charges are accumulated in this regime, but the doped organic layer is successively depleted. The thickness of the depletion layer varies with the distance from the source electrode; it is largest at the drain electrode, where potential difference between the channel and the gate electrode is largest and smallest at the source, where this potential difference is smallest. In between these two extrema, the depletion layer varies. If $V(x)$ denotes the potential in the channel at a distance x from the source (see Figure 24), the thickness of the depletion layer $d_{dep}(x)$ follows from equation 16 by replacing $V - V_{FB}$ by $V_{GS} - V(x) - V_{FB}$

$$d_{dep} = \frac{\epsilon_0 \epsilon_s}{C_{ins}} \left[\sqrt{1 + \frac{2C_{ins}^2 (V_{GS} - V(x) - V_{FB})}{eN_A \epsilon_0 \epsilon_s}} - 1 \right]. \quad (23)$$

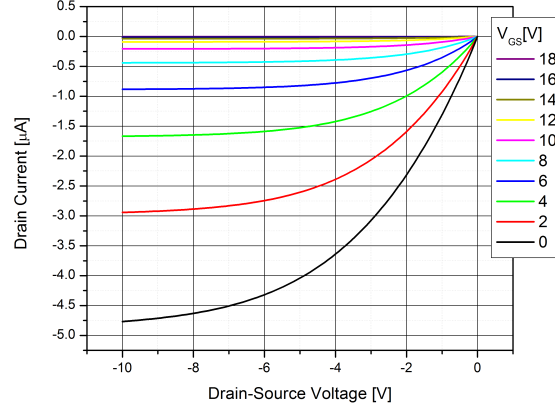


Figure 47: Output characteristic of the p-doped transistor sketched in 24 operated in the depletion regime. The p-doped channel is 8nm thick. Details of the device structure and preparation conditions are given in 21

The output characteristic of a doped pentacene transistor ($d_{dop} = 8$ nm) operated in the depletion region is shown in Figure 47. At low drain-source voltages and constant gate voltages the transistor shows a drain current I_D , which is proportional to the drain source voltage V_{DS} (linear regime). Thus, similar to undoped organic transistors, the transistors behave like a resistor, whose resistance is controlled by the gate voltage.

At a larger drain potential V_{DS} , the current saturates. This behavior can be understood by a full depletion of the doped layer at the drain contact, i.e. the channel is pinched off at the drain. Setting the depletion layer thickness d_{dep} to the thickness of the doped layer d_{dop} and replacing the potential $V - V_{FB}$ by $V_{GS} - V_{FB} - V_{DS}^{sat}$ (the potential difference between gate and channel at the drain) in equation 16, one finds the following condition for saturation:²⁶⁹

$$V_{DS}^{sat} = V_{GS} - V_{PO} \quad (24)$$

here, V_{PO} is the so called pinch-off voltage, defined as²⁶⁹

$$V_{PO} = V_{FB} + \frac{eN_A d_{org}^2}{2\epsilon_0 \epsilon_s} \left(1 + 2 \frac{d_i}{d_{org}} \frac{\epsilon_s}{\epsilon_i} \right) \quad (25)$$

with ϵ_i : relative permittivity of the insulator and d_i : thickness of insulator.

For drain-source voltages larger than the saturation voltage, the pinch-off point moves slightly toward the source.²⁸⁶ However, the potential at the pinch-off point remains at V_{DS}^{sat} . Furthermore, assuming that the channel length of the OFET is large (i.e. much larger than the movement of the pinch-off point) the electric field in the channel remains constant and hence the drain current saturates.

To completely turn the transistor off, the whole transistor channel has to be depleted. To reach a full depletion, the gate potential has to be large enough to deplete the channel at the source as well. Replacing $V - V_{FB}$ by $V_{GS} - V_{FB}$ in equation 16 leads to

$$V_{GS} = V_{PO}, \quad (26)$$

i.e. the pinch off voltage has to be applied to fully switch the transistor off.

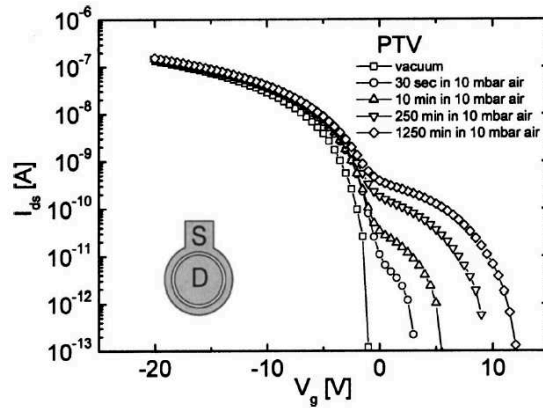


Figure 48: Increase in pinch-off voltage of polymer based transistors due to doping. The transistors consisting of poly(2,5-thienylene vinylene), PTV, were exposed to ambient air for increasing times, which led to an increase in p-doping due to oxygen. Reprinted from⁶⁸, with the permission of AIP Publishing.

The increase in the pinch off voltage by doping was studied systematically by Meijer et al.,⁶⁸ who exposed a circular polymer based OFET to ambient air for a well defined time. The result is shown in Figure 48. For longer exposure times, a higher voltage has to be applied to the gate to switch the channel off. Meijer et al. argue that this increase in the pinch-off voltage stems from p-doping due to oxygen. Indeed, using Equation 25 they were able to deduce the doping concentration, which ranges from 10^{-16} to 10^{-17} cm⁻³.

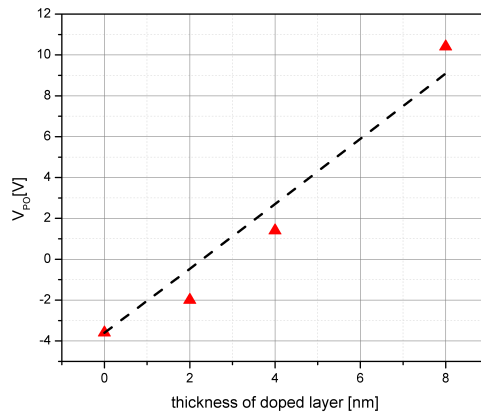


Figure 49: To avoid a decrease in ON/OFF ratio of doped organic transistors, the thickness of the doped layer d_{dop} can be limited to several nanometers. The influence of the thickness of the doped layer on the pinch-off voltage of p-doped pentacene transistors is shown above. Increasing the thickness to 8 nm leads to an increase in pinch-off voltage to more than 10 V.

The increase of the pinch-off voltage V_{PO} with doping concentration (Equation 25) often results in a decrease in the ON/OFF ratio of transistors.^{23,25,166,287–289} To solve this problem, either the doping concentration has to be very small,⁵¹ or the thickness of the doped layer d_{dop} has to be decreased to a couple of nanometers. The dependency of the pinch-off voltage on the thickness of the doped channel is shown in Figure 49. As seen in this Figure, not only does the threshold voltage V_{th} shift according to Equation 22 (Figure 44), but the pinch-off voltage increases as well from around -4 V for the intrinsic transistor to more than 10 V for the transistor with the 8 nm thick doped channel.

Instead of decreasing the doping concentration or the thickness of the doped layer, Lu

et al. doped an organic transistor consisting of a mixture of insulating polystyrene and P3HT as active layer.²⁸⁹ Indeed, they found that doping does not significantly increase the off current, but increases the on current. This phenomena was explained by the three-dimensional network of P3HT formed inside the polystyrene bulk. The network is sparse at the gate insulator/semiconductor interface, i.e. it is possible to deplete parts of the P3HT network at the insulator interface, leading to a complete off switching of the transistor.

A variant of the organic depletion transistors was reported in 2014 - the organic junction field-effect transistor (organic JFET).²⁹⁰ In the JFET the gate insulator is replaced by an organic p-i-n junction. Applying a reverse voltage to the junction leads to an increased depletion e.g. in the p-doped layer.⁵⁶ This increase in depletion layer thickness leads to a modulation of the conductance of the layer, which can be sensed by the source and drain contact.²⁹⁰

Organic electrochemical transistors (OETs) are another popular example of an organic transistor which is commonly operated in depletion mode.²⁹¹ OETs consist of source and drain contact and a highly doped organic semiconductors bridging the gap between the two electrodes. PEDOT:PSS (Section 3.1.3) is one of the most common choices as it combines a high electrical conductivity with a significant ionic mobility.

OETs are operated in ionic solutions, e.g. in salt water. If a positive voltage is applied between an electrode immersed into the solution and the transistor channel, mobile cations move into the organic semiconductor and de-dope the film. The conductivity of the film decreases, which is seen by a decrease in drain current.

It was shown by Bernards et al. that the density of cations entering and de-doping the organic semiconductor can be modeled by a capacitor, resulting in a straightforward electrical model of the transistor resembling the structure of standard depletion OFETs discussed above.²⁹² Indeed, the electrical behavior of OETs is very similar to doped OFETs. A representative output characteristic of an OET as reported by the Malliaras group is

shown in Figure 50.²⁹³

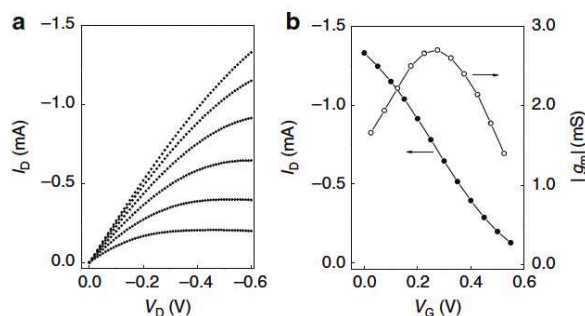


Figure 50: Output and transfer characteristic of an organic electrochemical transistor with a very large transconductance. Reprinted with permission from²⁹³. Copyright 2013 Nature Publishing Group.

OETs have the advantage that they convert ionic currents into electric signals, they operate in liquid environments, and some of the materials used - in particular PEDOT:PSS - are bio-compatible. Therefore, these transistors are often used in the field of organic bioelectronics.^{294,295} The field of OETs has been reviewed recently e.g. by Owens et al.,²⁹⁶ Kergoat et al.,²⁹⁷ and Ates,²⁹⁸ and the reader is referred to these reviews for a more detailed summary of the advances in the field of OETs.

5.2.3 Majority and Minority Charge Carriers in Doped Transistors

Doping adds a further dimension in the control of the electronic properties of organic semiconductors. Doping defines the majority charge carrier type; if the Fermi level is close to the HOMO state, holes are dominating, whereas electrons are dominating if the Fermi Level is close to the LUMO. This effect is seen for example in Figure 5 for the polycrystalline material ZnPc and Figure 6 for pentacene - adding more electrons to the material by doping leads to a shift of the Fermi Level toward the LUMO, whereas p-doping shifts the Fermi level toward the HOMO.

This control on the majority charge carrier type can be used to influence charge transport in organic semiconductors, which for long have been thought of as unipolar conductors. For example, Lee et al. showed that C_{60} based transistors can be operated as

p-type transistors as well, if they are heavily doped by the p-dopant MoO_3 ²⁰¹ (see Section 4.3 for a discussion of p-type doping of C_{60}).

Similarly, Naab et al. showed that TIPS-pentacene can form n-type transistors if doped by DMBI- derivatives (see Section 3.2.5) acting as n-type dopants.¹³³ Relying on their solution sheering method,²⁰⁸ very large n-type mobilities of up to $6.8 \text{ cm}^2/\text{Vs}$ were observed.

Doping can as well be used to control the balance in electron and hole currents in ambipolar organic transistors.^{213,299} In particular, as ambipolar transistors are often unwanted for complementary logic circuits, doping can be used to define the majority charge carrier type and to turn the ambipolar behavior on or off. The group of Noh et al. reported on selectively doping PCBM-type transistors.²¹³ The results are shown in Figure 39. Without doping, the transistors show a strong ambipolarity. Adding the n-dopant CsF to the layer of PCBM leads to a reduction in hole transport and increase in electron transport, whereas adding $\text{F}_4\text{-TCNQ}$ leads to a suppression of electron transport and increase in hole transport.

In all transistors discussed above, majority charge carriers are accumulated at the insulator/organic semiconductor interface. This type of operation is in marked contrast to the dominating inorganic transistor technology - inversion based MOSFETs. Here, a thin layer of minority charge carriers are accumulated at the insulator/semiconductor interface, e.g. holes are accumulated in an n-doped layer or electrons in a p-doped layer. The organic semiconductor is said to be inverted, i.e. transformed from a p-doped layer into an n-doped layer (or vice versa).

Very low off-currents and large ON/OFF ratios are a major advantage of inversion based FETs. For organic semiconductors, this suppression of off-state currents might prove to be beneficial in transistors with very low channel lengths, e.g. realized by vertically structuring the transistors.³⁰⁰ Indeed, it was shown recently that organic inversion FETs can be realized.^{21,212} The transistors consist of p-doped injection contacts to ensure an

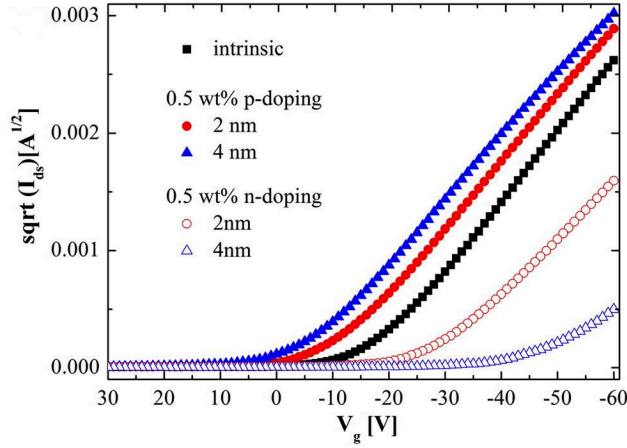


Figure 51: Doped top-gate organic transistors. Both, p- and n-doping lead to a shift in the threshold voltage. For n-doped channels, a p-minority channel has to be formed in the doped layer, i.e. the transistor is operated in inversion. Reprinted from²¹², with the permission of AIP Publishing.

efficient hole injection into the organic semiconductor pentacene. The channel region is doped by the n-dopant $W_2(hpp)_4$. Thus, to switch the transistor on, the n-doped channel has to be depleted first, before an inversion layer of holes can be formed at the insulator/organic interface. This effect was seen to result in a shift of the threshold voltage toward more negative voltages (see Figure 51).^{21,212}

6 Outlook and conclusions

Organic doping has become a key in the development of highly efficient optoelectronic devices³⁰¹ such as OLEDs⁴⁸ or OSCs.^{10,152} Compared to OLEDs and OSCs, doping of organic transistors is lagging behind and the vast majority of OFET publications is relying on intrinsic materials. However, in recent years the benefits of doping for OFETs have increasingly been accepted and a concise understanding of the influence of doping on the OFET performance has been developed.

This progress was made possible by a broad library of dopants developed in the past, which spans n- and p-type dopants and with various doping strength, sizes, and stability

in air. Doping has a strong influence on the structure of most polycrystalline materials as commonly used in thin film transistors. A wealth of studies have been reported that clarify the structure of doped organic high-mobility layers and several structural models were proposed. However, in most cases a reduction in crystallinity is still observed at high doping concentrations, leading to a reduction in charge carrier mobility and more efforts to design dopants that incorporate into the crystal lattice rather than disrupt the order in the matrix are needed.

Doping can be used at different positions in the device. The most common use of doping is to improve charge carrier injection and extraction at the source and drain contact. In fact, suppression of injection at the source electrode was identified as a major factor limiting the switching speed of best OFETs and realizing quasi-ohmic contacts by doping is a promising route to increase the transconductance and cutoff frequency of OFETs significantly.

Channel doping is another use of doping in OFETs. In the channel, doping will influence the flatband and threshold voltage and it will determine the pinch-off behavior of the channel. Thus, doping provides an important design parameter to control device performance. Furthermore, doping was reported to fill unwanted trap states in the device, which results in more stable devices and even increases the stability of n-type OFETs in air.

In summary, doping adds a new dimension to the design of highly-efficient OFETs. Considering the importance of doping in inorganic FETs, doping will in the future not only be used to improve conventional transistors, but it will enable new device concepts that are based on tuning and controlling the Fermi-Level position in the organic semiconductor.

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Björn Lüssem is an Assistant Professor of Physics at Kent State University. Before joining Kent State in 2014, he was a project leader at Materials Science Laboratory of Sony in Stuttgart (2006-2008) and group leader at the Institut für Angewandte Photophysik (IAPP), TU Dresden (2008-2013).

Lüssem authored more than 100 publications in international journals (h-index > 25). His research interests range from new design principles of organic devices, to charge transport in organic semiconductors and novel organic semiconductors with improved performance. His work has been awarded the VDE-Promotionspreis and the Günther-Leibfried-Preis.

Chang-Min Keum

Chang-Min Keum received his Ph.D. degree in electrical engineering from the Seoul National University (South Korea) in 2014 under the guidance of Prof. Sin-Doo Lee. During doctoral studies, his research topics were mainly on interface phenomena, device physics, and architecture of organic electronic devices such as organic field-effect transistors and vertical-structured organic light-emitting transistors. Since he joined the group of Prof. Lüssem at the Kent State University as a postdoctoral research associate in 2015, he has been working on enhancement of the light outcoupling efficiency in organic light-emitting diodes and development of vertical organic light-emitting transistors.

Daniel Kasemann

Dr. Daniel Kasemann, Doctor of Science (Dr. rer. nat., Physics), received his Ph.D. in 2011 for research on singlet-triplet-annihilation in OLEDs under extremely high current

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Ben Naab

Benjamin Naab received a Ph.D. in chemistry in 2015 from Stanford University in California for his thesis on molecular n-dopants for organic electronics. He is currently a chemist in Dow Electronic Materials of The Dow Chemical Company. His main research interests are in the field of molecular dopants for organic electronics and interconnect technologies for printed circuit boards. He is also interested in metallization chemistry for emerging dielectric materials.

Zhenan Bao

Zhenan Bao is a Professor of Chemical Engineering at Stanford University. Prior to joining Stanford in 2004, she was a Distinguished Member of Technical Staff in Bell Labs, Lucent Technologies from 1995-2004. She has over 400 refereed publications and over 60 US patents with a Google Scholar H-Index >120. She pioneered a number of design concepts for organic electronic materials. Her work has enabled flexible electronic circuits and displays. In her recent work, she has developed skin-inspired organic electronic materials, which resulted in unprecedented performance or functions in medical devices, energy storage and environmental applications.

Bao is an academic member of the National Academy of Engineering. She is a Fellow of MRS, ACS, AAAS, SPIE, ACS PMSE and ACS POLY. She served on the Board of Directors for MRS in 2003-2005 and as an Executive Committee Member for the Polymer

Materials Science and Engineering division of the American Chemical Society. She is an Associate Editor for Chemical Sciences.

Bao was selected as Nature's Ten people who mattered in 2015 for her work on artificial electronic skin. She was awarded the AIChE Andreas Acrivos Award for Professional Progress in Chemical Engineering in 2014, ACS Carl Marvel Creative Polymer Chemistry Award in 2013, ACS Cope Scholar Award in 2011, she was the recipient of the Royal Society of Chemistry Beilby Medal and Prize in 2009, the IUPAC Creativity in Applied Polymer Science Prize in 2008, American Chemical Society Team Innovation Award 2001, R&D 100 Award and R&D Magazine's Editors Choice of the "Best of the Best" new technology for 2001. She has been selected in 2002 by the American Chemical Society Women Chemists Committee as one of the twelve "Outstanding Young Woman Scientist who is expected to make a substantial impact in chemistry during this century". She was also selected by MIT Technology Review magazine in 2003 as one of the top 100 young innovators for this century.

Bao is a co-founder and on the Board of Directors for C3 Nano, a silicon-valley venture funded start-up commercializing flexible transparent electrodes.

Karl Leo

Karl Leo obtained the Diplomphysiker degree from the University of Freiburg in 1985, working with Adolf Goetzberger at the Fraunhofer-Institut für Solare Energiesysteme. In 1988, he obtained the PhD degree from the University of Stuttgart for a PhD thesis performed at the Max-Planck-Institut für Festkörperforschung in Stuttgart under supervision of Hans Queisser. From 1989 to 1991, he was postdoc at AT&T Bell Laboratories in Holmdel, NJ, U.S.A. From 1991 to 1993, he was with the Rheinisch-Westfälische Technische Hochschule (RWTH) in Aachen, Germany. Since 1993, he is full professor of optoelectronics at the Technische Universität Dresden. His main interests are novel semiconductor systems like semiconducting organic thin films; with special emphasis to

understand basics device principles and the optical response. His work was recognized by a number of awards, including: Otto-Hahn-Medaille (1989), Bennigsen-Förder-Preis (1991), Leibniz-Award (2002), award of the Berlin-Brandenburg Academy (2002), Manfred-von-Ardenne-Preis (2006), Zukunftspreis of the German president (2011), Rudolf-Jäckel-Prize (2012), and a Dr. techn. h.c. of the University of Southern Denmark (2013). He is cofounder of several companies, including Novaled GmbH and Heliatek GmbH.

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